

Experiment 25

B. Determination of the equilibrium constant for the iodine + iodide = triiodide reaction as a function of concentration.

1. Concentration of KI in original solution

$$\frac{2.00 \times 10^{-3} \text{ M}}{}$$

2. Concentration of I_2 in original solution

$$\frac{1.00 \times 10^{-4} \text{ M}}{}$$

3. Equilibrium expression for reaction

$$\frac{[I_3^-]}{[I_2][I^-]}$$

4. Results and calculations (Calculate the concentrations of iodine ($[I_2]_0$) and iodide ($[I^-]_0$) that would result from the dilutions if a reaction did not occur. Using the absorption measurement for triiodide (A) and its extinction coefficient, calculate the triiodide concentration ($[I_3^-]$) that results from the achievement of equilibrium, then calculate the final concentrations of iodine ($[I_2]$) and iodide ($[I^-]$) by subtraction. Finally, calculate the value of the equilibrium constant (K_{eq}).

$$A = \epsilon bc$$

#	mL I_2	mL KI	total vol. (mL)	$[I_2]_0$ (mol/L)	$[I^-]_0$ (mol/L)	A	Temp. ($^{\circ}C$)	$[I_3^-]$ (mol/L)	$[I_2]$ (mol/L)	$[I^-]$ (mol/L)	K_{eq}	H_2O mL
1	2.00	1.00	4.00	5.0×10^{-5}	5.0×10^{-3}	0.937	23 $^{\circ}$	3.09×10^{-5}	1.91×10^{-5}	4.69×10^{-3}	7.69×10^2	1.00
2	2.00	2.00	4.00	5.0×10^{-5}	1.0×10^{-2}	0.560	23	1.85×10^{-5}	3.2×10^{-5}	8.15×10^{-3}	4.71×10^2	N/A
3	3.00	1.00	4.00	7.5×10^{-5}	5.0×10^{-3}	0.456	23	1.20×10^{-5}				N/A
4	1.00	1.00	4.00	2.5×10^{-5}	5.0×10^{-3}	0.089	23	2.90×10^{-5}				2.00
5	1.00	3.00	4.00	2.5×10^{-5}	1.5×10^{-2}	0.204	23	7.06×10^{-6}				N/A

5. Average value of K_{eq}

Equilibrium Constant Determination Of Iodide

H Kauffman



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Table of Contents Equilibrium Constant Determination Of Iodide

1. Understanding the eBook Equilibrium Constant Determination Of Iodide
 - The Rise of Digital Reading Equilibrium Constant Determination Of Iodide
 - Advantages of eBooks Over Traditional Books
2. Identifying Equilibrium Constant Determination Of Iodide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Equilibrium Constant Determination Of Iodide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Equilibrium Constant Determination Of Iodide
 - Personalized Recommendations
 - Equilibrium Constant Determination Of Iodide User Reviews and Ratings
 - Equilibrium Constant Determination Of Iodide and Bestseller Lists

5. Accessing Equilibrium Constant Determination Of Iodide Free and Paid eBooks
 - Equilibrium Constant Determination Of Iodide Public Domain eBooks
 - Equilibrium Constant Determination Of Iodide eBook Subscription Services
 - Equilibrium Constant Determination Of Iodide Budget-Friendly Options
6. Navigating Equilibrium Constant Determination Of Iodide eBook Formats
 - ePub, PDF, MOBI, and More
 - Equilibrium Constant Determination Of Iodide Compatibility with Devices
 - Equilibrium Constant Determination Of Iodide Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Equilibrium Constant Determination Of Iodide
 - Highlighting and Note-Taking Equilibrium Constant Determination Of Iodide
 - Interactive Elements Equilibrium Constant Determination Of Iodide
8. Staying Engaged with Equilibrium Constant Determination Of Iodide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Equilibrium Constant Determination Of Iodide
9. Balancing eBooks and Physical Books Equilibrium Constant Determination Of Iodide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Equilibrium Constant Determination Of Iodide
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Equilibrium Constant Determination Of Iodide
 - Setting Reading Goals Equilibrium Constant Determination Of Iodide
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Equilibrium Constant Determination Of Iodide
 - Fact-Checking eBook Content of Equilibrium Constant Determination Of Iodide
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

-
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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