

EXERCISES

MULTIPLE-CHOICE QUESTIONS

8) Consider the following reaction:



Which one of the following statements is correct?

- A. Since very large K_c values are rare, this reaction is reversible.
- B. Changing the pressure will affect the value of K_c .
- C. In an open system, this reaction is effectively irreversible.
- D. Changing the temperature will not affect the value of K_c .

9) The following reaction, in which nitrogen monoxide, $\text{NO}(g)$, is converted to nitrogen dioxide, $\text{NO}_2(g)$, occurs at equilibrium:



In a given closed system, the equilibrium constant for this reaction is 0.20. At some time, the initial concentrations of $\text{NO}(g)$ at equilibrium is 0.2000 M. At this temperature, the initial concentration of $\text{N}_2\text{O}_4(g)$ at equilibrium is

- A. 0.0500 M
- B. 0.0010 M
- C. 0.0200 M
- D. 0.0020 M

10) Dissolved carbon dioxide, $\text{CO}_2(aq)$, reacts with water, $\text{H}_2\text{O}(l)$, to form, mathematically, $\text{H}_2\text{CO}_3(aq)$, $\text{H}_2\text{CO}_3^*(aq)$, a species that reacts with $\text{H}_2\text{O}(l)$ to form bicarbonate ions, $\text{HCO}_3^-(aq)$, and hydronium ions, $\text{H}_3\text{O}^+(aq)$, according to the equilibrium below:



It is known that increases in CO_2 at elevated depths in plants (in a closed and dark environment) can shift these two biochemical equilibria. In this scenario, what change is made in these systems, which one of the following changes is most likely to result in an increase in the concentration of $\text{H}_3\text{O}^+(aq)$?

- A. Change the volume.
- B. Double the mass $\text{CO}_2(aq)$.
- C. Add a few drops of phenolphthalein indicator.
- D. Add a few drops of 0.1 M sodium hydroxide, $\text{NaOH}(aq)$, solution.

Equilibrium Test 2 Multiple Choice Answers

Ronald Pirayoff



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