

Honors Chemistry
Equilibrium Worksheet

Equilibrium: ICE box practice problems

1. In a 10.0L vessel at 1000K, 0.250 mol $\text{SO}_2(\text{g})$ and 0.200 mol $\text{O}_2(\text{g})$ react to form 0.162 mol $\text{SO}_3(\text{g})$ at equilibrium. What is the K_c at 1000K for this reaction?

Balanced equation: $\text{SO}_2(\text{g}) + \text{O}_2(\text{g}) \leftrightarrow \text{SO}_3(\text{g})$

	SO_2	+	O_2	\leftrightarrow	SO_3
I					
C					
E					

2. An aqueous solution of 0.00325M Fe^{3+} was reacted with 0.0020M SCN^- at 25°C. Calculate the value of the equilibrium constant for this reaction when the equilibrium concentration of FeSCN^{2+} was determined to be 0.000667M

	Fe^{3+}	+	SCN^-	\leftrightarrow	FeSCN^{2+}
I					
C					
E					

3. At a certain temperature, 2.0 mole NH_3 is introduced into a 2.0 L container, and the NH_3 partially dissociates into by the following reaction:



At equilibrium, 1.0 mol NH_3 remains, what is the value of K for this reaction?

	$2\text{NH}_3(\text{g})$	\leftrightarrow	$\text{N}_2(\text{g})$	+	$3\text{H}_2(\text{g})$
I					
C					
E					

Equilibrium Ice Box Practice Problems

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Equilibrium Ice Box Practice Problems:

Decoding **Equilibrium Ice Box Practice Problems**: Revealing the Captivating Potential of Verbal Expression

In a period characterized by interconnectedness and an insatiable thirst for knowledge, the captivating potential of verbal expression has emerged as a formidable force. Its ability to evoke sentiments, stimulate introspection, and incite profound transformations is genuinely awe-inspiring. Within the pages of "**Equilibrium Ice Box Practice Problems**," a mesmerizing literary creation penned by a celebrated wordsmith, readers attempt an enlightening odyssey, unraveling the intricate significance of language and its enduring impact on our lives. In this appraisal, we shall explore the book's central themes, evaluate its distinctive writing style, and gauge its pervasive influence on the hearts and minds of its readership.

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