

EXP. 1 MEASUREMENT OF AN EQUILIBRIUM CONSTANT

POST-LAB REPORT WORKSHEETS

Concentration of standard $\text{Na}_2\text{S}_2\text{O}_3$ solution: 0.01105

Temperature of the equilibrium mixture: 21.1 °C

Part 1 – Titration of the Aqueous Layer

You must have 3 titrations.

Volume of Aqueous Aliquots:	1	2	3
Final Burette Reading:	<u>19.61</u>	<u>36.51</u>	<u>41.68</u>
Initial Burette Reading:	<u>10.54</u>	<u>20.53</u>	<u>30.21</u>
Volume of Aliquot:	<u>9.07</u>	<u>9.98</u>	<u>11.47</u>
Titration with Sodium Thiosulfate:			
Final Burette Reading:	<u>35.98</u>	<u>37.18</u>	<u>38.93</u>
Initial Burette Reading:	<u>3.34</u>	<u>4.88</u>	<u>4.81</u>
Volume of Titer:	<u>32.64</u>	<u>32.30</u>	<u>34.12</u>
Total Concentration of Reducible Iodine [$\text{I}_2 + \text{I}_3^-$] in Aqueous Phase:	<u>0.0199</u>	<u>0.0179</u>	<u>0.01644</u>

Equilibrium Constant Lab Report Discussion

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