

## Practice Shifting Equilibrium Problems

Name \_\_\_\_\_

Per \_\_\_\_\_

1-16: For each of the following systems, tell whether the reaction will shift to the left (reactant side), right (product side), or will experience no shift. State your reasoning. In each case, all reactants and products are gases.



- a) pressure is decreased
- b) temperature is raised
- c) D is removed from the system



- a) pressure is increased
- b) temperature is lowered
- c) D is added to the system



- a) pressure is increased
- b) temperature is lowered
- c) C is added to the system



- a) pressure is decreased
- b) temperature is raised



- a) pressure is increased
- b) temperature is raised
- c) A is added to the system



- a) pressure is decreased
- b) temperature is decreased
- c) A is added to the system



- a) pressure is decreased
- b) temperature is decreased
- c) B is removed



- a) pressure is increased
- b) temperature is raised
- c) volume of the container is increased



- a) pressure is increased
- b) temperature is lowered
- c) volume of the container is increased



- a) B is added to the system
- b) temperature is raised
- c) volume of the container is increased



- a) volume of the container is increased
- b) the system is heated



- The reaction is endothermic
- a) pressure is reduced
- b) the system is cooled



- The reaction is exothermic
- a) pressure is decreased
- b) heat is removed from the system
- c) C is removed from the system



- The system is exothermic
- a) volume of the container is increased
- b) heat is added to the system



- The system is exothermic
- a) pressure is lowered
- b) temperature is lowered



- The system is endothermic
- a) pressure is increased
- b) temperature is lowered
- c) B is added to the system

# Equilibrium Shift Practice Problems

**JG Myers**



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