

Equilibrium Practice Problems

1. Write the equilibrium expression for each of the following reactions:



$$K = \frac{[\text{NH}_3]^2}{[\text{N}_2][\text{H}_2]^3}$$



$$K = \frac{[\text{ICl}]^2}{[\text{I}_2][\text{Cl}_2]}$$



$$K = \frac{[\text{NO}][\text{O}_2]^{1/2}}{[\text{NO}_2]}$$

2. The dissociation of acetic acid, CH_3COOH , has an equilibrium constant at 25°C of 1.8×10^{-5} . The reaction is



If the equilibrium concentration of CH_3COOH is 0.46 moles in 0.500 L of water and that of CH_3COO^- is 8.1×10^{-3} moles in the same 0.500 L, calculate $[\text{H}^+]$ for the reaction.

$$K = \frac{[\text{CH}_3\text{COO}^-][\text{H}^+]}{[\text{CH}_3\text{COOH}]} \quad \text{Plug in known values and solve.}$$

$$1.8 \times 10^{-5} = \frac{[8.1 \times 10^{-3} \text{ moles}/0.500 \text{ L}][\text{H}^+]}{[0.46 \text{ moles}/0.500 \text{ L}]}$$

$$[\text{H}^+] = 1.0 \times 10^{-3} \text{ M}$$

3. Indicate the effect of a catalyst, pressure, temperature and concentration on each of the following on:

	catalyst	pressure	temperature	concentration
a. speed	presence increase	direct relationship	direct relationship	direct relationship
b. pos ⁿ	nothing	based on moles	based on enthalpy of rxn	inverse relationship

Equilibrium Practice Problems Problem And Answers

Ying-Ying Zheng



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Table of Contents Equilibrium Practice Problems Problem And Answers

1. Understanding the eBook Equilibrium Practice Problems Problem And Answers
 - The Rise of Digital Reading Equilibrium Practice Problems Problem And Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Equilibrium Practice Problems Problem And Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Equilibrium Practice Problems Problem And Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Equilibrium Practice Problems Problem And Answers
 - Personalized Recommendations
 - Equilibrium Practice Problems Problem And Answers User Reviews and Ratings

- Equilibrium Practice Problems Problem And Answers and Bestseller Lists
- 5. Accessing Equilibrium Practice Problems Problem And Answers Free and Paid eBooks
 - Equilibrium Practice Problems Problem And Answers Public Domain eBooks
 - Equilibrium Practice Problems Problem And Answers eBook Subscription Services
 - Equilibrium Practice Problems Problem And Answers Budget-Friendly Options
- 6. Navigating Equilibrium Practice Problems Problem And Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Equilibrium Practice Problems Problem And Answers Compatibility with Devices
 - Equilibrium Practice Problems Problem And Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Equilibrium Practice Problems Problem And Answers
 - Highlighting and Note-Taking Equilibrium Practice Problems Problem And Answers
 - Interactive Elements Equilibrium Practice Problems Problem And Answers
- 8. Staying Engaged with Equilibrium Practice Problems Problem And Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Equilibrium Practice Problems Problem And Answers
- 9. Balancing eBooks and Physical Books Equilibrium Practice Problems Problem And Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Equilibrium Practice Problems Problem And Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Equilibrium Practice Problems Problem And Answers
 - Setting Reading Goals Equilibrium Practice Problems Problem And Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Equilibrium Practice Problems Problem And Answers
 - Fact-Checking eBook Content of Equilibrium Practice Problems Problem And Answers
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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