

Balancing Equations Worksheet

- 1) $\underline{\quad} \text{Na}_3\text{PO}_4 + \underline{\quad} \text{KOH} \rightarrow \underline{\quad} \text{NaOH} + \underline{\quad} \text{K}_3\text{PO}_4$
- 2) $\underline{\quad} \text{MgF}_2 + \underline{\quad} \text{Li}_2\text{CO}_3 \rightarrow \underline{\quad} \text{MgCO}_3 + \underline{\quad} \text{LiF}$
- 3) $\underline{\quad} \text{P}_4 + \underline{\quad} \text{O}_2 \rightarrow \underline{\quad} \text{P}_2\text{O}_3$
- 4) $\underline{\quad} \text{RbNO}_3 + \underline{\quad} \text{BeF}_2 \rightarrow \underline{\quad} \text{Be}(\text{NO}_3)_2 + \underline{\quad} \text{RbF}$
- 5) $\underline{\quad} \text{AgNO}_3 + \underline{\quad} \text{Cu} \rightarrow \underline{\quad} \text{Cu}(\text{NO}_3)_2 + \underline{\quad} \text{Ag}$
- 6) $\underline{\quad} \text{CF}_4 + \underline{\quad} \text{Br}_2 \rightarrow \underline{\quad} \text{CBr}_4 + \underline{\quad} \text{F}_2$
- 7) $\underline{\quad} \text{HCN} + \underline{\quad} \text{CuSO}_4 \rightarrow \underline{\quad} \text{H}_2\text{SO}_4 + \underline{\quad} \text{Cu}(\text{CN})_2$
- 8) $\underline{\quad} \text{GaF}_3 + \underline{\quad} \text{Cs} \rightarrow \underline{\quad} \text{CsF} + \underline{\quad} \text{Ga}$
- 9) $\underline{\quad} \text{BaS} + \underline{\quad} \text{PtF}_2 \rightarrow \underline{\quad} \text{BaF}_2 + \underline{\quad} \text{PtS}$
- 10) $\underline{\quad} \text{N}_2 + \underline{\quad} \text{H}_2 \rightarrow \underline{\quad} \text{NH}_3$
- 11) $\underline{\quad} \text{NaF} + \underline{\quad} \text{Br}_2 \rightarrow \underline{\quad} \text{NaBr} + \underline{\quad} \text{F}_2$
- 12) $\underline{\quad} \text{Pb}(\text{OH})_2 + \underline{\quad} \text{HCl} \rightarrow \underline{\quad} \text{H}_2\text{O} + \underline{\quad} \text{PbCl}_2$
- 13) $\underline{\quad} \text{AlBr}_3 + \underline{\quad} \text{K}_2\text{SO}_4 \rightarrow \underline{\quad} \text{KBr} + \underline{\quad} \text{Al}_2(\text{SO}_4)_3$
- 14) $\underline{\quad} \text{CH}_4 + \underline{\quad} \text{O}_2 \rightarrow \underline{\quad} \text{CO}_2 + \underline{\quad} \text{H}_2\text{O}$
- 15) $\underline{\quad} \text{Na}_3\text{PO}_4 + \underline{\quad} \text{CaCl}_2 \rightarrow \underline{\quad} \text{NaCl} + \underline{\quad} \text{Ca}_3(\text{PO}_4)_2$
- 16) $\underline{\quad} \text{K} + \underline{\quad} \text{Cl}_2 \rightarrow \underline{\quad} \text{KCl}$
- 17) $\underline{\quad} \text{Al} + \underline{\quad} \text{HCl} \rightarrow \underline{\quad} \text{H}_2 + \underline{\quad} \text{AlCl}_3$
- 18) $\underline{\quad} \text{N}_2 + \underline{\quad} \text{F}_2 \rightarrow \underline{\quad} \text{NF}_3$
- 19) $\underline{\quad} \text{SO}_2 + \underline{\quad} \text{Li}_2\text{Se} \rightarrow \underline{\quad} \text{SSe}_2 + \underline{\quad} \text{Li}_2\text{O}$
- 20) $\underline{\quad} \text{NH}_3 + \underline{\quad} \text{H}_2\text{SO}_4 \rightarrow \underline{\quad} (\text{NH}_4)_2\text{SO}_4$

Equation Writing And Balancing 8

David Baud

A decorative graphic element consisting of a light blue horizontal bar with a rounded right end, and a red-to-white gradient semi-circle on the right side of the bar.

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