

1. K_1 and K_2 are the velocity constants of forward and backward reactions



The equilibrium constant k of the reaction is

Sol. (C)

2. Which of the following reactions will be favoured at low pressure?



Sol. (C)

3. For the reaction, $A + B \longrightarrow C + D$, if active mass of A is kept constant and active mass of B is tripled, the rate of reaction will become

(A) three times

(B) six times

(C) eight times

(D) nine times

Sol. (D)

4. The rate of $\frac{K_p}{K_c}$ for the reaction:



(A) 1

(B) RT

(C) $(RT)^{1/2}$

(D) $(RT)^{-1/2}$

Sol. (D)

5. Consider the reaction,



in a closed container at equilibrium. What would be the effect of addition of $CaCO_3$ on the equilibrium concentration of CO_2 ?

(A) Increases

(B) Unpredictable

(C) Decreases

(D) Remains unaffected

Sol. (D)

6. For the reaction, $PCl_5(g) \rightleftharpoons PCl_3(g) + Cl_2(g)$, the forward reaction at constant temperature is favoured by

(A) introducing an inert gas at constant volume

(B) introducing $PCl_5(g)$ at constant volume

(C) introducing $PCl_3(g)$ at constant volume

(D) introducing $Cl_2(g)$ at constant volume

Sol. (C)

7. A liquid is in equilibrium with its vapour at its boiling point. On an average, the molecules in the two phases have equal

(A) intermolecular forces

(B) potential energy

(C) kinetic energy

(D) both kinetic energy and potential energy

Sol. (C)

Multiple Choice Questions In Chemical Equilibrium

Marco Cascella



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