

ANSWER KEY

ATOMS AND ELECTRONS STUDY GUIDE

Learning Objectives: (1) (10-1) P. 41 - Explain a new scientific theory of atomic-scale phenomena (atomic theory) by describing changes in the atomic model over time, including those changes that are associated with experimental evidence. (10-1) P. 42 - Explain the relative physical and chemical behavior of atoms by describing the structure of atoms in terms of positive nucleus, orbiting electrons, and the relative spacing between protons and neutrons in the nucleus. (10-1) P. 42 - Explain properties of atoms and their changes in the periodic table in terms of arrangements of their electrons, related to P. 42, as well as the periodic table's organization in terms of fundamental (atomic) properties (atomic number, mass number, atomic symbol).

QUESTION BANK

Directions: Match the atomic theory statement to its correct conclusion.

Statement	Conclusion	Correct Match
1. Electrons are made of atoms.	A. Electrons are made of atoms.	A
2. Electrons are made of atoms.	B. Electrons are made of atoms.	B
3. Electrons are made of atoms.	C. Electrons are made of atoms.	C
4. Electrons are made of atoms.	D. Electrons are made of atoms.	D
5. Electrons are made of atoms.	E. Electrons are made of atoms.	E
6. Electrons are made of atoms.	F. Electrons are made of atoms.	F

Directions: Use the periodic table to answer the following questions:

1. **Atomic Number** = **P. 42, 43** - The number of protons equals the number of electrons.
2. **Atomic Mass** = **P. 42, 43** - The number of protons and neutrons add equal the **Atomic Mass**.
3. **P. 42, 43** - **Atomic Number** = **Atomic Mass** - **Atomic Number** = **Atomic Mass** - **Atomic Number** = **Atomic Mass**.

Energy And Electrons Study Guide Answers

Jiyuan Zhang



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