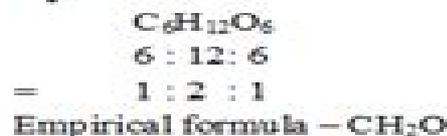
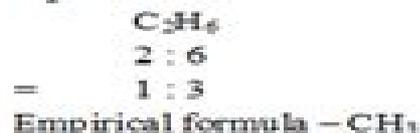


Empirical formula: It shows the simplest whole number ratio between the atoms of the elements in the compound.

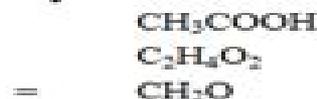
Example 1:



Example 2:



Example 3:



Practice:

1. C_4H_8
2. $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_3$
3. N_2O_4

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Calculate the empirical formula of the compounds having the following percentage composition:

Example 1:

C-40%
H-6.7%
O-53.3%

Sol:

Element	%	% ÷ At. Wt.	Simplest ratio	simplest whole number ratio
Carbon	40	$40 \div 12 = 3.33$	$\frac{3.33}{3.33} = 1$	1 C
Hydrogen	6.7	$6.7 \div 1$	$\frac{6.7}{3.33} = 2$	2 H
Oxygen	53.3	$53.3 \div 16$	$\frac{3.33}{3.33} = 1$	1 O
Empirical Formula		= CH₂O		

Empirical Formula Practice

Michal Rosen-Zvi

A decorative graphic consisting of a semi-transparent red circle with a white center, positioned on the right side of the slide, partially overlapping the light blue bar.

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