

# Empirical and Molecular Formula Practice

3 pages + 3-page Answer KEY + Periodic Table

## EMPIRICAL AND MOLECULAR PROBLEMS

Chemistry

1 The percent composition of a certain gas is nitrogen (80.0%) and oxygen (20.0%). Calculate the empirical formula.

Step 1. Convert % to g.

80 g N

20 g O

Step 2. Convert g to mol.

Step 3. Divide by the smallest mole value.

Step 4. Multiply by smallest number that will make a whole number.

2 The compound in problem 1 has a molar mass of 100 g/mol. Calculate the molecular formula.

Step 1. Calculate the molar mass of the empirical formula.

Step 2. Divide molecular molar mass by empirical molar mass.

Step 3. Multiply empirical formula by the answer in Step 2.

Name: \_\_\_\_\_ Date: \_\_\_\_\_

3 A compound is composed of 82.7% nitrogen and 17.3% hydrogen. What is the empirical formula of this compound?

4 The compound in problem 3 has a molar mass of 100 g/mol. Calculate the molecular formula.

5 Hydrazine, a high energy fuel, is 75% nitrogen and 25% hydrogen. Calculate the empirical formula.

6 Hydrazine's molar mass is 62 g/mol. Calculate the molecular formula.

Name: \_\_\_\_\_ Date: \_\_\_\_\_

7 A compound, composed of hydrogen (5.9%) and oxygen (94.1%), has a molar mass of 54 g/mol. What is the empirical formula of this compound?

8 What is the molecular formula of the compound in problem 7?

9 A compound with a known molar mass of 28 g/mol contains carbon (85.7%) and hydrogen (14.3%). What is the empirical formula?

10 What is the molecular formula of the compound in problem 9?

ANSWER KEY INCLUDED

# Empirical And Molecular Formulas Answer Key

**N Noddings**



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*Basic Chemistry* William S. Seese, Guido H. Daub, 1977

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### **Table of Contents Empirical And Molecular Formulas Answer Key**

1. Understanding the eBook Empirical And Molecular Formulas Answer Key
  - The Rise of Digital Reading Empirical And Molecular Formulas Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Empirical And Molecular Formulas Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Empirical And Molecular Formulas Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Empirical And Molecular Formulas Answer Key
  - Personalized Recommendations

- Empirical And Molecular Formulas Answer Key User Reviews and Ratings
- Empirical And Molecular Formulas Answer Key and Bestseller Lists
- 5. Accessing Empirical And Molecular Formulas Answer Key Free and Paid eBooks
  - Empirical And Molecular Formulas Answer Key Public Domain eBooks
  - Empirical And Molecular Formulas Answer Key eBook Subscription Services
  - Empirical And Molecular Formulas Answer Key Budget-Friendly Options
- 6. Navigating Empirical And Molecular Formulas Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Empirical And Molecular Formulas Answer Key Compatibility with Devices
  - Empirical And Molecular Formulas Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Empirical And Molecular Formulas Answer Key
  - Highlighting and Note-Taking Empirical And Molecular Formulas Answer Key
  - Interactive Elements Empirical And Molecular Formulas Answer Key
- 8. Staying Engaged with Empirical And Molecular Formulas Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Empirical And Molecular Formulas Answer Key
- 9. Balancing eBooks and Physical Books Empirical And Molecular Formulas Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Empirical And Molecular Formulas Answer Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Empirical And Molecular Formulas Answer Key
  - Setting Reading Goals Empirical And Molecular Formulas Answer Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Empirical And Molecular Formulas Answer Key
  - Fact-Checking eBook Content of Empirical And Molecular Formulas Answer Key

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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