

Questions Lab Course Inorganic Chemistry 2019-2020

1. Why must we not use concentrated sulphuric acid for the synthesis of iron(II) sulfate?
2. Which role plays the iron nail in solutions containing Fe^{2+} ions? Write the corresponding reactions.
3. An iron nail will be dissolved by diluted sulfuric acid. What will happen with a copper nail? What will be result when you treat both iron and copper nails with concentrated sulfuric acid?
4. It is recommended to store iron(II) salts in closed containers. Why?
5. Describe the structure of alum. Which kind of interatomic and intermolecular forces can you recognize in the structure?
6. How can we explain the existence and formation of double sulfates?
7. Are double sulfates stable in aqueous solution? Give reasons for your answer.
8. The Fe^{2+} in Mohr's salt is significantly more stable as in other Fe^{2+} salts. Why?
9. Justify why the X-ray pattern of alum is clearly different from that of Mohr's salt although both compounds have identical atoms and similar chemical composition.
10. White copper(I) chloride is easily oxidized in air forming a green double salt with formula $\text{CuCl}_2 \cdot 3\text{Cu}(\text{OH})_2$. Explain why (tip: use the corresponding normal potentials).
11. Balance the different reactions for obtaining chlorine gas starting from hydrochloric acid and a) potassium permanganate; b) manganese(IV) oxide and c) potassium dichromate. Calculate its thermodynamic spontaneity.
12. How can you calculate the experimental yield of the synthesis of a metal halide that is seen to be unstable in humid air?
13. During formation of iron(III) chloride:
 - a) indicate oxidizing and reducing agent
 - b) Explain though being a redox process, you must not use reduction potentials for calculating its spontaneity.
14. Explain why carbon tetrachloride a) does not suffer hydrolysis and b) is immiscible with water.
15. Arrange the following metal halides from lower to higher melting point. Give reasons for your answer: MF_3 , MCl_3 , MI_3
16. Arrange the following metal halides by increasing ionicity of bond. Give reasons for your answer: NaCl , ClF , FeCl_3 , FeCl_2 , F_2 , SiCl_4
17. Draw the structure of the ferricacate complex. Analyse the complex concerning central atom, type and number of ligands, coordination number, geometry and hybridization of central atom. What kind of isomerism does the complex present?

Inorganic Lab Vaiva Questions

R Sandford



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