

The purpose of this dry lab is to familiarize you with the shapes of small molecules in 3 dimensions.
 The molecular shape is determined by the number of atoms directly bonded to the central atom as well as the number of lone pairs on the central atom. The below table can be used to determine molecular shape.

On paper, we draw 3-D molecules in a specific way.
 Where **A** and **X** represent any element, and **A** is the central atom.

1. Molecular Shapes

Regions of Electron Density around Central Atom	Number of bonds on the central atom	Number of lone pairs on the central atom	Molecular Shape	Generic Structure	Examples
1	1 bond	Any number	Linear	$A-X$	$H-\ddot{Cl}:$
2	2 bonds	0 lone pairs	Linear	$X-A-X$	$H-C\equiv C-H$
3	3 bonds	0 lone pairs	Trigonal Planar		
	2 bonds	1 lone pair	Bent		
4	4 bonds	0 lone pairs	Tetrahedral		
	3 bonds	1 lone pair	Pyramidal		
	2 bonds	2 lone pairs	Bent		

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