

Molar Mass and Mole Calculations Worksheet

1. Calculate the molar mass for each of the following compounds:

- | | |
|--|---|
| a. NaCl = 58.44 g/mol | e. $(\text{NH}_4)_3\text{PO}_4 = 149.12 \text{ g/mol}$ |
| b. $\text{Fe}_2\text{O}_3 = 159.7 \text{ g/mol}$ | f. $\text{Ca}_3(\text{PO}_4)_2 = 310.18 \text{ g/mol}$ |
| c. $\text{MgCO}_3 = 84.32 \text{ g/mol}$ | g. $\text{C}_6\text{H}_{12}\text{O}_6 = 180.18 \text{ g/mol}$ |
| d. $\text{Ba}(\text{OH})_2 = 171.35 \text{ g/mol}$ | h. $\text{C}_2\text{H}_5\text{N} = 79.11 \text{ g/mol}$ |

2. Calculate the number of moles in each of the following masses:

- a. 45.0 g of acetic acid, $\text{CH}_3\text{COOH} = 0.75 \text{ mol}$
- b. 7.04 g of lead (II) nitrate = **0.02 mol**

3. Calculate the mass in grams of each of the following amounts:

- a. 3.00 mol of selenium oxybromide, $\text{SeOBr}_2 = 764.31 \text{ g}$
 $M = 78.97 + 16 + 2(79.9) = 254.77 \text{ g/mol} \quad 254.77 \times 3 =$
- b. 488 mol of calcium carbonate = **48,843.92 g**
 $M = 40.08 + 12.01 + 3(16) = 100.09 \text{ g/mol} \quad 100.09 \times 488 =$

4. Calculate the number of molecules or formula units in each of the following amounts:

- a. 4.27 mol of tungsten(VI) oxide, $\text{WO}_3 = 2.57 \times 10^{24}$
 $4.27 \times 6.02 \times 10^{23} =$
- b. 0.989 mol of nitrobenzene $\text{C}_6\text{H}_5\text{NO}_2 = 5.95 \times 10^{23}$
 $0.989 \times 6.02 \times 10^{23} =$

5. Calculate the mass of each of the following quantities:

- a. 8.39×10^{23} molecules of fluorine = **52.96 g**
 $8.39 \times 10^{23} / 6.02 \times 10^{23} = 1.39 \text{ mol} \quad 1.39 \times 38 =$
 $M = 2(19) = 38 \text{ g/mol}$
- b. 6.82×10^{24} formula units of beryllium sulfate = **1,190.33 g**
 $M = 9.01 + 32.06 + 4(16) = 105.07 \quad 105.07 \times 11.33 =$
 $6.82 \times 10^{24} / 6.02 \times 10^{23} = 11.33$

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Table of Contents Mole Calculations Study Guide Answer Key

1. Understanding the eBook Mole Calculations Study Guide Answer Key
 - The Rise of Digital Reading Mole Calculations Study Guide Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Mole Calculations Study Guide Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mole Calculations Study Guide Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mole Calculations Study Guide Answer Key
 - Personalized Recommendations
 - Mole Calculations Study Guide Answer Key User Reviews and Ratings
 - Mole Calculations Study Guide Answer Key and Bestseller Lists
5. Accessing Mole Calculations Study Guide Answer Key Free and Paid eBooks
 - Mole Calculations Study Guide Answer Key Public Domain eBooks
 - Mole Calculations Study Guide Answer Key eBook Subscription Services
 - Mole Calculations Study Guide Answer Key Budget-Friendly Options
6. Navigating Mole Calculations Study Guide Answer Key eBook Formats
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 - Mole Calculations Study Guide Answer Key Compatibility with Devices
 - Mole Calculations Study Guide Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mole Calculations Study Guide Answer Key
 - Highlighting and Note-Taking Mole Calculations Study Guide Answer Key
 - Interactive Elements Mole Calculations Study Guide Answer Key

8. Staying Engaged with Mole Calculations Study Guide Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mole Calculations Study Guide Answer Key
9. Balancing eBooks and Physical Books Mole Calculations Study Guide Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mole Calculations Study Guide Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Mole Calculations Study Guide Answer Key
 - Setting Reading Goals Mole Calculations Study Guide Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Mole Calculations Study Guide Answer Key
 - Fact-Checking eBook Content of Mole Calculations Study Guide Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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