

Honors Chemistry
Equilibrium Worksheet

Equilibrium: ICE box practice problems

1. In a 10.0L vessel at 1000K, 0.250 mol $\text{SO}_2(\text{g})$ and 0.200 mol $\text{O}_2(\text{g})$ react to form 0.162 mol $\text{SO}_3(\text{g})$ at equilibrium. What is the K_c at 1000K for this reaction?

Balanced equation: $\text{SO}_2(\text{g}) + \text{O}_2(\text{g}) \leftrightarrow \text{SO}_3(\text{g})$

	SO_2	+	O_2	\leftrightarrow	SO_3
I					
C					
E					

2. An aqueous solution of 0.00325M Fe^{3+} was reacted with 0.0020M SCN^- at 25°C. Calculate the value of the equilibrium constant for this reaction when the equilibrium concentration of FeSCN^{2+} was determined to be 0.000667M

	Fe^{3+}	+	SCN^-	\leftrightarrow	FeSCN^{2+}
I					
C					
E					

3. At a certain temperature, 2.0 mole NH_3 is introduced into a 2.0 L container, and the NH_3 partially dissociates into by the following reaction:



At equilibrium, 1.0 mol NH_3 remains, what is the value of K for this reaction?

	$2\text{NH}_3(\text{g})$	\leftrightarrow	$\text{N}_2(\text{g})$	+	$3\text{H}_2(\text{g})$
I					
C					
E					

Equilibrium Ice Box Practice Problems

R Bogdan



Equilibrium Ice Box Practice Problems:

A-level Physics Challenging Practice Questions (Yellowreef) Thomas Bond, Chris Hughes, 2013-11-14 first to completely cover all question types since 1996 with answer keys first to expose all trick questions provides full set of step by step solution approaches available separately provides an easy path to final A distinction grade Complete edition and concise edition eBooks available

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