

Equilibrium Practice Test

Multiple Choice

Identify the choice that best completes the statement or answers the question.

- _____ 1. A reaction in which products can react to re-form reactants is
a. reforming. c. buffered.
b. reversible. d. impossible.
- _____ 2. At equilibrium,
a. all reactions have ceased.
b. only the forward reaction continues.
c. only the reverse reaction continues.
d. both the forward and reverse reactions continue.
- _____ 3. What symbol in a chemical equation indicates equilibrium?
a. \rightarrow c. \rightleftharpoons
b. \leftarrow d. None of the above
- _____ 4. What two processes are at equilibrium in a saturated sugar solution?
a. evaporation and condensation c. decomposition and synthesis
b. dissolving and crystallization d. ionization and recombination
- _____ 5. What symbol represents the equilibrium constant?
a. k c. c
b. K d. R
- _____ 6. A very high value of K indicates that
a. equilibrium is reached slowly. c. reactants are favored.
b. products are favored. d. equilibrium has been reached.
- _____ 7. The value of K for a system
a. can be calculated from the molar masses of products and reactants.
b. can be calculated from the enthalpies of the forward and reverse reactions.
c. can be calculated from the chemical properties of products and reactants.
d. must be measured by experiment.
- _____ 8. In the equation $K = \frac{[W][X]}{[Y][Z]}$, what represents the concentrations of the reactants?
a. $[Y]$ and $[Z]$ c. $\frac{[W][X]}{[Y][Z]}$
b. $[W]$ and $[X]$ d. $K = \frac{[W][X]}{[Y][Z]}$
- _____ 9. An equilibrium mixture of SO_2 , O_2 , and SO_3 gases at 1500 K is determined to consist of 0.344 mol/L SO_2 , 0.172 mol/L O_2 , and 0.56 mol/L SO_3 . What is the equilibrium constant for the system at this temperature? The balanced equation for this reaction is
$$2\text{SO}_2(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{SO}_3(\text{g})$$

a. 0.41 c. 6.7
b. 2.8 d. 15

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Table of Contents Kinetics And Equilibrium Practice Test

1. Understanding the eBook Kinetics And Equilibrium Practice Test
 - The Rise of Digital Reading Kinetics And Equilibrium Practice Test
 - Advantages of eBooks Over Traditional Books
2. Identifying Kinetics And Equilibrium Practice Test
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Kinetics And Equilibrium Practice Test
 - User-Friendly Interface
4. Exploring eBook Recommendations from Kinetics And Equilibrium Practice Test
 - Personalized Recommendations
 - Kinetics And Equilibrium Practice Test User Reviews and Ratings
 - Kinetics And Equilibrium Practice Test and Bestseller Lists

5. Accessing Kinetics And Equilibrium Practice Test Free and Paid eBooks
 - Kinetics And Equilibrium Practice Test Public Domain eBooks
 - Kinetics And Equilibrium Practice Test eBook Subscription Services
 - Kinetics And Equilibrium Practice Test Budget-Friendly Options
6. Navigating Kinetics And Equilibrium Practice Test eBook Formats
 - ePub, PDF, MOBI, and More
 - Kinetics And Equilibrium Practice Test Compatibility with Devices
 - Kinetics And Equilibrium Practice Test Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Kinetics And Equilibrium Practice Test
 - Highlighting and Note-Taking Kinetics And Equilibrium Practice Test
 - Interactive Elements Kinetics And Equilibrium Practice Test
8. Staying Engaged with Kinetics And Equilibrium Practice Test
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Kinetics And Equilibrium Practice Test
9. Balancing eBooks and Physical Books Kinetics And Equilibrium Practice Test
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Kinetics And Equilibrium Practice Test
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Kinetics And Equilibrium Practice Test
 - Setting Reading Goals Kinetics And Equilibrium Practice Test
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Kinetics And Equilibrium Practice Test
 - Fact-Checking eBook Content of Kinetics And Equilibrium Practice Test
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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