

ANSWERS TO UNIT I : REACTION KINETICS

1. Rate of consumption = $\frac{5.0 \text{ g}}{150 \text{ s}} \times \frac{60 \text{ s}}{1 \text{ min}} = \frac{2.0 \text{ g}}{\text{min}}$
2. Time = $45.0 \text{ g} \times \frac{1 \text{ min}}{2.35 \text{ g}} = 19.1 \text{ min}$
3. Volume of $\text{O}_2 = 7.50 \text{ min} \times \frac{32.5 \text{ mL}}{\text{min}} = 244 \text{ mL}$
4. (a) OK; moles/second = amount/time
 (b) Not acceptable; this is a "rate" but not "amount/time"
 (c) OK; moles/litre is an amount and (moles/litre)/second = amount/time
 (d) Not acceptable; this is a density or concentration
 (e) OK; millilitres/hour = amount/time
 (f) OK; grams/minute = amount/time
5. Rate of consumption of $\text{O}_2 = \frac{1.34 \text{ mol H}_2\text{O}}{\text{min}} \times \frac{1 \text{ mol O}_2}{2 \text{ mol H}_2\text{O}} = \frac{0.67 \text{ mol O}_2}{\text{min}}$
6. (a) The mass is decreasing because $\text{CO}_2(\text{g})$ is being lost from the solution.
 (b) Slope = -0.006
 (c) i) unit of rise = g ii) unit of run = s iii) units of slope = g/s
 (d) Expected units for rate = g/s
 (e) The rate of the reaction = the slope of the line = -0.006 g/s (a loss of 0.006 g every second)
7. If the reaction occurred in a closed container, no mass would be lost (conservation of mass). If the $\text{NO}_2(\text{g})$ was allowed to escape, the mass decrease would be due to the loss of $\text{NO}_2(\text{g})$, NOT the loss of $\text{Cu}(\text{s})$. [Incidentally, $\text{NO}_2(\text{g})$ is quite soluble in water and not all of the NO_2 produced would escape.]
8. (a) Measure the rate at which the intensity of the blue colour in the solution increases.
 (b) Measure the change in mass of $\text{Cu}(\text{s})$ and calculate the rate as the mass of Cu used divided by the time required for the reaction; or measure the change in mass of $\text{Ag}(\text{s})$ and calculate the rate as the mass of Ag produced divided by the reaction time; or measure the increase in temperature (the reaction is exothermic) and calculate the rate as the temperature change divided by reaction time.
9. (a) The reaction produces two moles of gas for every two moles of gas used up. Hence the number of moles of gas in the container doesn't change, and the gas pressure doesn't change as the reaction proceeds. (Recall "Avogadro's Hypothesis" from Chem 11: equal volumes of gases at the same temperature and pressure contain the same number of molecules or moles.)
 (b) Rate = $\frac{1.2 \text{ g HCl}}{120 \text{ s}} \times \frac{1 \text{ mol HCl}}{36.5 \text{ g HCl}} = 2.7 \times 10^{-4} \text{ mol HCl/s}$
 (c) Rate (of HCl) = $\frac{0.200 \text{ L H}_2}{1 \text{ min}} \times \frac{2 \text{ L HCl}}{1 \text{ L H}_2} = 0.400 \text{ L HCl/min}$
10. Reaction (b) is fastest: a single H^+ is removed from NH_4^+ and added to CO_3^{2-} .
 Reaction (a) is moderately slow: 6 bonds have to be broken and made. An $\text{O}-\text{O}$ bond and 2 $\text{O}-\text{H}$ bonds have to be broken; 2 $\text{O}-\text{H}$ bonds and a second bond in $\text{O}=\text{O}$ must be made.
 Reaction (c) is very slow: many bonds have to be broken and made. Incidentally ... gasoline does not undergo combustion at ROOM TEMPERATURE; a mixture of gasoline and oxygen can exist indefinitely at room temperature. Of course, if you add HEAT (in the form of a spark or match) then BOOM!
11. Since nothing was changed, other than exchanging one reactant (Li) for another (K), the reaction rate change was due to "the nature of the reactants".

Kinetics Of A Reaction Lab Answers

AW Rasmussen



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