

CHAPTER SIX

- 6.1 a.** AB: Heat liquid - $-V \approx \text{constant}$
BC: Evaporate liquid - $-P$ increases, system remains at point on vapor - liquid equilibrium curve as long as some liquid is present. $T = 100^\circ\text{C}$.
CD: Heat vapor - $-T$ increases, V increases.

- b.** Point B: Neglect the variation of the density of liquid water with temperature, so $\rho = 1.00 \text{ g/mL}$ and $V_b = \underline{10 \text{ mL}}$.

Point C: H_2O (v, 100°C)

$$n = \frac{10 \text{ mL}}{1 \text{ mL}} \left| \frac{1.00 \text{ g}}{\text{mL}} \right| \left| \frac{1 \text{ mol}}{18.02 \text{ g}} \right| = 0.555 \text{ mol}$$

$$P_c V_c = nRT_c \Rightarrow V_c = \frac{nRT_c}{P_c} = \frac{0.555 \text{ mol}}{1 \text{ atm}} \left| \frac{0.08206 \text{ L} \cdot \text{atm}}{\text{mol} \cdot \text{K}} \right| \left| \frac{373 \text{ K}}{1} \right| = \underline{17 \text{ L}}$$

- 6.2 a.** $P_{\text{total}} = \underline{243 \text{ mm Hg}}$. Since liquid is still present, the pressure and temperature must lie on the vapor-liquid equilibrium curve, where by definition the pressure is the vapor pressure of the species at the system temperature.

- b.** Assuming ideal gas behavior for the vapor,

$$m(\text{vapor}) = \frac{(3.000 - 0.010) \text{ L}}{(30 + 273.2) \text{ K}} \left| \frac{\text{mol} \cdot \text{K}}{0.08206 \text{ L} \cdot \text{atm}} \right| \left| \frac{243 \text{ mm Hg}}{760 \text{ mm Hg}} \right| \left| \frac{1 \text{ atm}}{1} \right| \left| \frac{119.39 \text{ g}}{\text{mol}} \right| = 4.59 \text{ g}$$

$$m(\text{liquid}) = \frac{10 \text{ mL}}{1 \text{ mL}} \left| \frac{1.489 \text{ g}}{\text{mL}} \right| = 14.89 \text{ g}$$

$$m_{\text{total}} = m(\text{vapor}) + m(\text{liquid}) = \underline{19.5 \text{ g}}$$

$$x_{\text{vapor}} = \frac{4.59}{19.48} = \underline{0.235 \text{ g vapor / g total}}$$

- 6.3 a.** $\log_{10} p^* = 7.09808 - \frac{1238.71}{45 + 217} = 2.370 \Rightarrow p^* = 10^{2.370} = \underline{234.5 \text{ mm Hg}}$

$$\text{b. } \ln p^* = -\frac{\Delta \hat{H}_v}{R} \frac{1}{T} + B \Rightarrow -\frac{\Delta \hat{H}_v}{R} = \frac{\ln(p_2^* / p_1^*)}{\frac{1}{T_2} - \frac{1}{T_1}} = \frac{\ln(760 / 118.3)}{\frac{1}{(77.0 - 273.2)\text{K}} - \frac{1}{(29.5 - 273.2)\text{K}}} = -4151 \text{ K}$$

$$B = \ln(p_1^*) + \frac{\Delta \hat{H}_v / R}{T_1} = \ln(118.3) + \frac{4151 \text{ K}}{(29.5 + 273.2)\text{K}} = 18.49$$

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