

Name \_\_\_\_\_

Section \_\_\_\_\_

**Experiment 1****Advance Study Assignment: Densities of Solids and Liquids**

The advance study assignments in this laboratory manual are designed to assist you in making the calculations required in the experiment you will be doing. We do this by furnishing you with sample data and showing in some detail how that data can be used to obtain the desired results. In the advance study assignments we will often include the guiding principles as well as the specific relationships to be employed. If you work through the steps in each calculation by yourself, you should have no difficulty when you are called upon to make the necessary calculations on the basis of the data you obtain in the laboratory.

**1. Finding the volume of a flask.**

A student obtained a clean, dry glass-stoppered flask. She weighed the flask and stopper on an analytical balance and found the total mass to be 33.695 g. She then filled the flask with water and obtained a mass for the full stoppered flask of 64.356 g. From these data, and the fact that at the temperature of the laboratory the density of water was 0.9978 g/cm<sup>3</sup>, find the volume of the stoppered flask.

- a. First we need to obtain the mass of the water in the flask. This is found by recognizing that the mass of a sample is equal to the sum of the masses of its parts. For the filled stoppered flask:

Mass of filled stoppered flask = mass of empty stoppered flask + mass of water,  
so mass of water = mass of filled flask - mass of empty flask

Mass of water = \_\_\_\_\_ g - \_\_\_\_\_ g = \_\_\_\_\_ g

Many mass and volume measurements in chemistry are made by the method used in 1a. This method is called measuring by difference, and is a very useful one.

- b. The density of a pure substance is equal to its mass divided by its volume:

$$\text{Density} = \frac{\text{mass}}{\text{volume}} \quad \text{or} \quad \text{volume} = \frac{\text{mass}}{\text{density}}$$

The volume of the flask is equal to the volume of the water it contains. Since we know the mass and density of the water, we can find its volume and that of the flask. Make the necessary calculation.

Volume of water = volume of flask = \_\_\_\_\_ cm<sup>3</sup>

**2. Finding the density of an unknown liquid.**

Having obtained the volume of the flask, the student emptied the flask, dried it, and filled it with an unknown whose density she wished to determine. The mass of the stoppered flask when completely filled with liquid was 58.540 g. Find the density of the liquid.

- a. First we need to find the mass of the liquid by measuring by difference:

Mass of liquid = \_\_\_\_\_ g - \_\_\_\_\_ g = \_\_\_\_\_ g

(continued on following page)

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**M Tight**



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