

## 5

## INTERPRETING GRAPHICS

Use with Section 5.3

$$c = 3.00 \times 10^8 \text{ m/s}$$

$$E = hc/\lambda$$

$$c = 3.00 \times 10^8 \text{ m/s}$$

$$h = 6.626 \times 10^{-34} \text{ J}\cdot\text{s}$$



$$E_n = -13.6 \text{ eV} / n^2$$

Figure 1 The emission spectrum and orbit-transition diagram for hydrogen.

Table 1

Transition	$E(\text{J})$	$\nu(\text{s}^{-1})$	$\lambda(\text{nm})$	Type of Radiation
$n = 5 \rightarrow n = 4$	$2.66 \times 10^{-18}$	$4.01 \times 10^{14}$	$7.47 \times 10^{-7}$	Infrared
$n = 5 \rightarrow n = 3$	$7.57 \times 10^{-18}$	$1.14 \times 10^{15}$	$2.63 \times 10^{-7}$	IR <sub>A</sub>
$n = 5 \rightarrow n = 2$	$1.82 \times 10^{-17}$	$2.75 \times 10^{15}$	$1.09 \times 10^{-6}$	IR <sub>B</sub>
$n = 5 \rightarrow n = 1$	$4.84 \times 10^{-17}$	$7.31 \times 10^{15}$	$4.10 \times 10^{-7}$	visible
$n = 4 \rightarrow n = 3$	$2.02 \times 10^{-18}$	$3.20 \times 10^{14}$	$9.37 \times 10^{-7}$	ultraviolet
$n = 4 \rightarrow n = 2$	$4.90 \times 10^{-18}$	$7.41 \times 10^{14}$	$4.05 \times 10^{-7}$	IR <sub>A</sub>
$n = 4 \rightarrow n = 1$	$1.55 \times 10^{-17}$	$2.34 \times 10^{15}$	$1.29 \times 10^{-6}$	IR <sub>B</sub>
$n = 3 \rightarrow n = 2$	$4.56 \times 10^{-18}$	$6.88 \times 10^{14}$	$4.36 \times 10^{-7}$	visible
$n = 3 \rightarrow n = 1$	$2.09 \times 10^{-17}$	$3.15 \times 10^{15}$	$9.52 \times 10^{-8}$	ultraviolet
$n = 2 \rightarrow n = 1$	$1.64 \times 10^{-17}$	$2.47 \times 10^{15}$	$1.22 \times 10^{-7}$	IR <sub>A</sub>
$n = 2 \rightarrow n = 1$	$4.09 \times 10^{-18}$	$6.17 \times 10^{14}$	$4.86 \times 10^{-7}$	visible
$n = 2 \rightarrow n = 1$	$2.04 \times 10^{-17}$	$3.10 \times 10^{15}$	$9.67 \times 10^{-8}$	ultraviolet
$n = 1 \rightarrow n = 1$	$3.00 \times 10^{-18}$	$4.57 \times 10^{14}$	$6.56 \times 10^{-7}$	visible
$n = 1 \rightarrow n = 1$	$1.94 \times 10^{-17}$	$2.93 \times 10^{15}$	$1.02 \times 10^{-7}$	ultraviolet
$n = 1 \rightarrow n = 1$	$1.64 \times 10^{-17}$	$2.48 \times 10^{15}$	$1.21 \times 10^{-7}$	ultraviolet

# Interpreting Graphics Chemistry Worksheet

**RC Schank**

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