

ANSWER KEY

ATOMS AND ELECTRONS STUDY GUIDE

Learning Objectives: (1) (10-1) P. 61 - Explain a new scientific theory of atomic-scale interactions (atomic theory) by describing changes in the atomic model over time, including those changes that are associated with experimental evidence. (10-1) P. 62 - Explain the relative historical dates when scientists began thinking about atoms, the structure of atoms, and how scientists determine atomic structure. Identify particles (protons, neutrons, and electrons) and their masses (in kg) within the atom. (10-1) P. 62 - Explain properties of atoms and their masses (in kg) within atoms (in kg) and their composition of their structure. (10-1) P. 62 - Explain the importance of atoms in the fundamental (atomic) structure of matter (atoms, molecules, ions, etc.).

ANSWER KEY

Directions: Match the atomic theory selected on the correct conditions.

Model/Year	Scientist(s)	Atomic Theory / Atomic Model Theory
(A)	Democritus (460 BC)	(A) Everything is made of atoms. Atoms cannot be created, divided or destroyed.
(B)	Bohr (1913)	(B) When Electrons move between energy levels, they emit or absorb light energy.
(C)	Rutherford (1911)	(C) Atoms are indivisible particles.
(D)	Thomson (1897)	(D) Atoms have a positively charged nucleus with electrons orbiting around it.
(E)	Modern (1920s)	(E) Electrons move in discrete orbits.
(F)	Avogadro (1808)	(F) Atoms are made of very small particles called molecules.

Directions: Use the periodic table to answer the following questions:

- (1) **10-1** (10-1) P. 62 - The number of protons equals the number of electrons.
- (2) **10-1** (10-1) P. 62 - The number of protons will equal the **atomic number**.
- (3) **10-1** (10-1) P. 62 - The number of electrons in a neutral atom is equal to the number of protons in the nucleus.

Electrons In Atoms Answer Key Study Guide

David R. Klein



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