Quiz 13-2

Name ANSWER KEY Class _____ Date _____

Assessment

Ions in Aqueous Solutions and Colligative Properties

Section Quiz: Colligative Properties of Solutions

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

B 1. Why is freezing-point depression a colligative property?
- a. It is inversely proportional to the number of particles in a solution.
- b. It is directly proportional to the number of particles in a solution.
- c. It depends on the properties of an electrolyte in a solvent.
- d. None of the above

B 2. Compared with a 0.01 M $C_6H_{12}O_6$ solution, a 0.01 M KCl solution has
- a. the same freezing-point depression.
- b. about twice the freezing-point depression.
- c. the same freezing-point elevation.
- d. about six times the freezing-point elevation.

$C_6H_{12}O_6$ $i=1$ KCl $i=2$

D 3. Compared with a 1.00 M NaI solution, a 1.00 M Na_2SO_4 solution has
- a. the same boiling-point elevation.
- b. about twice the boiling-point elevation.
- c. a boiling-point elevation about two-thirds as high.
- d. a boiling-point elevation about 1.5 times as high.

NaI $i=2$ $Na_2(SO_4)$ $i=3$

B 4. When a nonvolatile solute dissolves in a solvent, the vapor pressure of the solvent
- a. increases.
- b. decreases.
- c. stays the same.
- d. changes depending on the solvent used.

Vapor Pressure lowering (Colligative Property)

C 5. Which of the following compounds would be most effective in lowering the melting point of ice on roads?
- a. $CaCl_2$
- b. NaCl
- c. K_3PO_4
- d. K_2SO_4

$CaCl_{2(s)} \xrightarrow{H_2O} Ca^{+2}_{(aq)} + 2Cl^-_{(aq)}$ $i=3$

$NaCl_{(s)} \xrightarrow{H_2O} Na^+_{(aq)} + Cl^-_{(aq)}$ $i=2$

$K_3(PO_4)_{(s)} \xrightarrow{H_2O} 3K^+_{(aq)} + (PO_4)^{-3}_{(aq)}$ $i=4$

$K_2(SO_4)_{(s)} \xrightarrow{H_2O} 2K^+_{(aq)} + (SO_4)^{-2}$ $i=3$

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