

Quiz
132

Assessment

Ions in Aqueous Solutions and Colligative Properties

Section Quiz: Colligative Properties of Solutions

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

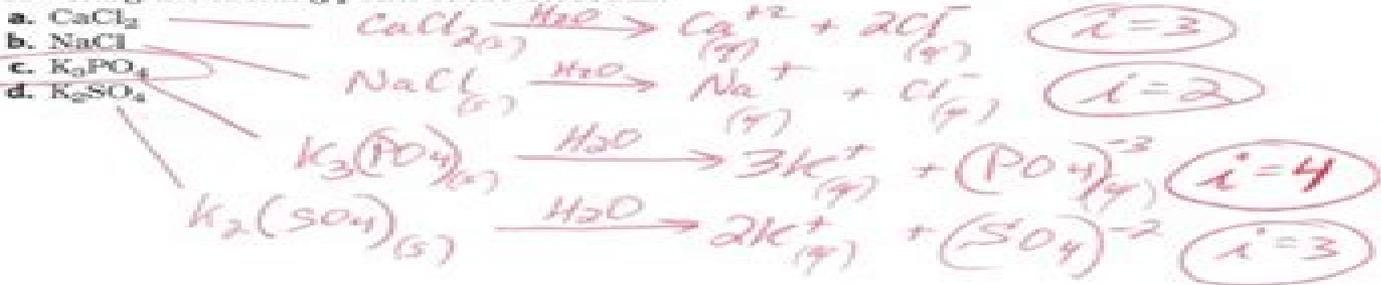
- B 1. Why is freezing-point depression a colligative property?
- a. It is inversely proportional to the number of particles in a solution.
 - b. It is directly proportional to the number of particles in a solution.
 - c. It depends on the properties of an electrolyte in a solvent.
 - d. None of the above

- B 2. Compared with a 0.01 M $C_6H_{12}O_6$ solution, a 0.01 M KCl solution has
- a. the same freezing-point depression.
 - b. about twice the freezing-point depression.
 - c. the same freezing-point elevation.
 - d. about six times the freezing-point elevation.
- Handwritten notes: $C_6H_{12}O_6$ (i=1), KCl (i=2)*

- D 3. Compared with a 1.00 M NaI solution, a 1.00 M Na_2SO_4 solution has
- a. the same boiling-point elevation.
 - b. about twice the boiling-point elevation.
 - c. a boiling-point elevation about two-thirds as high.
 - d. a boiling-point elevation about 1.5 times as high.
- Handwritten notes: NaI i=2, $Na_2(SO_4)$ i=3*

- B 4. When a nonvolatile solute dissolves in a solvent, the vapor pressure of the solvent
- a. increases.
 - b. decreases.
 - c. stays the same.
 - d. changes depending on the solvent used.
- Handwritten note: Vapor Pressure Lowering (Colligative Property)*

- C 5. Which of the following compounds would be most effective in lowering the melting point of ice on roads?



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BM King



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