

9 • Bonding & Molecular Structure

LEWIS STRUCTURE PRACTICE - 1

Ca^{2+} $[\text{Ca}]^{2+}$	NO_3^- $\leftrightarrow \left[\begin{array}{c} \text{O} \\ \\ \text{O} - \text{N} = \text{O} \\ \\ \text{O} \end{array} \right]^-$	PCl_5 $\begin{array}{c} \text{Cl} & & \text{Cl} \\ & & \\ \text{Cl} - \text{P} - \text{Cl} \\ \\ \text{Cl} \end{array}$
C_2H_2 $\text{H} : \text{C} : : : \text{C} : \text{H}$	S^{2-} $\left[: \ddot{\text{S}} : \right]^{2-}$	NH_3 $\text{H} : \text{N} : \text{H}$ $ $ H
SO_4^{2-} $\left[\begin{array}{c} \text{O} & & \text{O} \\ & & \\ \text{O} - \text{S} - \text{O} \\ & & \\ \text{O} & & \text{O} \end{array} \right]^{2-}$	SO_2 $\text{O} = \text{S} = \text{O} \leftrightarrow \text{O} = \text{S} - \text{O}$	CaH_2 $\text{H} : \text{Ca} : \text{H}$



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LEWIS STRUCTURE PRACTICE - 2

CH_4 $\begin{array}{c} \text{H} \\ \\ \text{H} - \text{C} - \text{H} \\ \\ \text{H} \end{array}$	HCN $\text{H} : \text{C} : : : \text{N} :$	K^+ $[\text{K}]^+$
PO_4^{3-} $\left[\begin{array}{c} \text{O} & & \text{O} \\ & & \\ \text{O} - \text{P} - \text{O} \\ & & \\ \text{O} & & \text{O} \end{array} \right]^{3-}$	O_2 $\text{O} = \text{O} \leftrightarrow \text{O} = \text{O}$	SF_6 $\begin{array}{c} \text{F} \\ \\ \text{F} - \text{S} - \text{F} \\ & & \\ \text{F} & & \text{F} \end{array}$
Br^- $\left[: \ddot{\text{Br}} : \right]^-$	AlH_3 $\text{H} : \text{Al} : \text{H}$ $ $ H	NH_4^+ $\left[\begin{array}{c} \text{H} \\ \\ \text{H} - \text{N} - \text{H} \\ \\ \text{H} \end{array} \right]^+$

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