

9 • Bonding & Molecular Structure

LEWIS STRUCTURE PRACTICE - 1

| | | |
|--|--|--|
| Ca^{2+} $[\text{Ca}]^{2+}$ | NO_2^- $\leftrightarrow \left[\begin{array}{c} \text{O} \\ \\ \text{O} - \text{N} = \text{O} \\ \\ \text{O} \end{array} \right]^-$ | PCl_5 $\begin{array}{c} (\text{Cl}) \\ \\ (\text{Cl}) - \text{P} - (\text{Cl}) \\ \\ (\text{Cl}) \end{array}$ |
| C_2H_2 $\text{H} : \text{C} :: \text{C} : \text{H}$ | S^{2-} $\left[\text{S} \right]^{2-}$ | NH_3 $\text{H} : \text{N} : \text{H}$ $ $ H |
| SO_4^{2-} $\left[\begin{array}{c} \text{O} \\ \\ \text{O} - \text{S} - \text{O} \\ \\ \text{O} \end{array} \right]^{2-}$ | SO_2 $\text{O} = \text{S} = \text{O} \leftrightarrow \text{O} = \text{S} - \text{O}$ | CaH_2 $\text{H} : \text{Ca} : \text{H}$ |



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LEWIS STRUCTURE PRACTICE - 2

| | | |
|--|--|--|
| CH_4 $\begin{array}{c} \text{H} \\ \\ \text{H} - \text{C} - \text{H} \\ \\ \text{H} \end{array}$ | HCN $\text{H} : \text{C} :: \text{N} :$ | K^+ $[\text{K}]^+$ |
| PO_4^{3-} $\left[\begin{array}{c} \text{O} \\ \\ \text{O} - \text{P} - \text{O} \\ \\ \text{O} \end{array} \right]^{3-}$ | O_2 $\text{O} = \text{O} \leftrightarrow \text{O} = \text{O}$ | SF_6 $\begin{array}{c} (\text{F}) \\ \\ (\text{F}) - \text{S} - (\text{F}) \\ \\ (\text{F}) \end{array}$ |
| Br^- $\left[\text{Br} \right]^-$ | AlH_3 $\text{H} : \text{Al} : \text{H}$ $ $ H | NH_4^+ $\left[\begin{array}{c} \text{H} \\ \\ \text{H} - \text{N} - \text{H} \\ \\ \text{H} \end{array} \right]^+$ |

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