



# CHEMISTRY

## CHEAT SHEET



# LEWIS STRUCTURES

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Lewis Structures are simplified representations of the valence electrons in molecules or compounds. They are important in understanding how atoms bond together. Follow the directions below to learn how to draw them.

**Directions**

1. Write the chemical formula for the molecule or compound.
2. Count the total number of valence electrons for the molecule or compound. Add the number of valence electrons for each atom in the formula.
3. Write the skeletal structure of the molecule.
4. Distribute the valence electrons around the atoms, starting with the most electronegative atom. Use all the valence electrons. If there are not enough electrons to give every atom an octet, then the molecule is a radical.

**Check/Tip**

1. All atoms must have a full outer shell of electrons.
2. The total number of valence electrons must be the same as the number of valence electrons in the original formula.

| Formulas for Lewis Structures with 1-4 Valence Electrons |  |                     |  |
|--|--|---------------------|--|
| 1 Valence Electrons                                      |  | 2 Valence Electrons |  |
| 2 Valence Electrons                                      |  | 3 Valence Electrons |  |
| 3 Valence Electrons                                      |  | 4 Valence Electrons |  |
| 4 Valence Electrons                                      |  | 5 Valence Electrons |  |

**Examples:** Draw the Lewis Structures for Magnesium (Mg), Boron (B), and Neon (Ne)

Magnesium has 2 valence electrons



Boron has 3 valence electrons



Neon has 8 valence electrons



Remember: Neutral molecules must have a total charge of zero. Lone pairs are 2 electrons in each pair.

# Lewis Structure Chemistry Study Guide

**CO Houle**



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