

3. Write the Nernst equation for the following cell reaction :



How will the E_{cell} be affected when concentration of

(i) Cu^{2+} ions is increased and

(ii) Zn^{2+} ions is increased ?

$1 \times 2 = 2$

SECTION - B

4. Following ions of 3d-transition series are given :



(Atomic number : Ti = 22, V = 23, Cr = 24, Mn = 25)

Identify the ion which is

(i) most stable in aqueous solution.

(ii) a strong oxidising agent.

(iii) colourless in aqueous solution.

Give suitable reason in each.

$1 \times 3 = 3$

5. (a) (i) On the basis of crystal field theory, write the electronic configuration for d^4 ion if $\Delta_0 < P$.

(ii) Using valence bond theory, predict the hybridization and magnetic character of $[\text{Ni}(\text{CN})_4]^{2-}$.

(Atomic number of Ni = 28)

(iii) Write the formula of the following complex using IUPAC norms :

Dichloridobis (ethane-1,2-diamine) cobalt (III)

$1 \times 3 = 3$

OR

(b) When a co-ordination compound $\text{NiCl}_2 \cdot 6\text{H}_2\text{O}$ mixed with AgNO_3 , 2 moles of AgCl are precipitated per mole of the compound. Write

(i) Structural formula of the complex.

(ii) Secondary valency of 'Ni' in the complex.

(iii) IUPAC name of the complex.

$1 \times 3 = 3$



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