

Experiment 5 Advance Study Assignment: Identification of a Compound by Mass Relationships

Name _____ Section _____

1. A student attempted to identify an unknown compound by the method used in this experiment. She found that when she heated a sample weighing 0.5032 g, the mass dropped to 0.2663 g. When the product was converted to a chloride, the mass went back up, to 0.3747 g.

a. Is the sample a carbonate? _____ Yes / no (Circle one)

Please provide your reasoning below.

b. What are the two compounds that might be in the unknown?

_____ or _____

c. Write the balanced chemical equation for the overall reaction that occurs when each of these two original compounds is converted to a chloride. If the compound is a hydrogen carbonate, use the sum of Reactions 1 and 2. If the sample is a carbonate, use Reaction 2. Write the equation for a sodium salt and then for a potassium salt.

d. How many moles of the chloride salt would be produced from one mole of the original compound?

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