

General Chemistry Semester 2 Study Guide

Mr. Pitzer - Fall 2018

1st Period: Monday, June 11th, 8:00am - 10:00am

2nd Period: Wednesday, June 13th, 11:00am - 1:00pm

3rd Period: Thursday, June 14th, 11:00am - 1:00pm

Use your resources wisely throughout the semester. You may not get to the end of the course by a college of other students that are less active than themselves. Maximizing your learning means the use of resources that you are provided to provide, but I expect a high level of self-motivation. Please review your notes, text, and journal for additional insight.

Make yourself a learning schedule of the semester. Consider your weaknesses as a student. Think of your school and how you can improve the learning experience. It is likely, at some stage that you will have your own study, test schedule. Use the completed schedule to identify whether you have met or will meet the required educational objectives.

Be flexible, creative, and diligent in your learning. Supplement, improve, and enhance your learning.

Study Guide: Review List

Make sure you have the following resources necessary available and that you have access to any materials that you need.

1. Periodic Table
2. Chemical Properties Table
3. Molar Masses Table
4. Thermodynamic Tables
5. Learning your ability to recognize and use chemical formulas, chemical reactions, nomenclature
6. Unit conversions (know these well) for use often:
 - a. 10^3 mg = 1 g
 - b. 1000 L = 1 m³
 - c. 1000 g = 1 kg
7. Chemical properties (know your nomenclature)

Final Exam: Reading Challenge

Review your notes, text, and learning resources and resources available throughout the semester. Make sure you have all appropriate resources available.

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| <ol style="list-style-type: none">1. Molecular Structure of Gases (pp. 101 - 110)<ol style="list-style-type: none">a. Understanding Avogadro's lawb. Molecular weightc. Molecular formula2. Kinetics (pp. 111 - 119, 121 - 124)<ol style="list-style-type: none">a. Rate lawsb. Half-life3. Thermodynamics (pp. 125 - 134)<ol style="list-style-type: none">a. Enthalpy (Calorimetry)b. Specific Heat4. Equilibrium (pp. 135 - 144) | <ol style="list-style-type: none">5. Kinetics (pp. 111 - 119)<ol style="list-style-type: none">a. Rate lawsb. Half-lifec. Molecular weightd. Molecular formula6. Kinetics (pp. 111 - 119)<ol style="list-style-type: none">a. Rate lawsb. Half-lifec. Molecular weightd. Molecular formula7. Kinetics (pp. 111 - 119)<ol style="list-style-type: none">a. Rate lawsb. Half-lifec. Molecular weightd. Molecular formula8. Kinetics (pp. 111 - 119)<ol style="list-style-type: none">a. Rate lawsb. Half-lifec. Molecular weightd. Molecular formula |
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