

(2)

- (c) Calculate the entropy change of a system containing a perfect gas when 1.00 mol of the gas doubles its volume at any temperature. 2

3. (a) For the reaction, $N_2 + 3H_2 \rightarrow 2NH_3$ the rate is expressed as

$$r = -\frac{d[N_2]}{dt} = -\frac{1}{3} \frac{d[H_2]}{dt} = \frac{1}{2} \frac{d[NH_3]}{dt}$$

- Under what conditions these expressions are valid? 1

- (b) The rate law of a homogeneous reaction $A \rightarrow P$ was found to be $r = k[A]^{1.38}$. What does the rate law indicate? 1

- (c) What are zeolites? Give one example of a reaction catalyzed by zeolite catalyst. 2

4. Answer any two of the following : 3×2=6

- (a) Calculate the work done when 50 g of iron (molar mass = 55.85 g mol^{-1}) reacts with hydrochloric acid to produce hydrogen gas in an open beaker at 25°C .

Engineering 2012 Course Chemistry 7 May 2014
Question Paper

VM Jensen



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