

## CHAPTER 2

# ACIDS, BASES AND SALTS



✓ You have learnt about acids, bases and salts as well as some of their properties in previous classes. Have you ever wondered why when a red ant or a wasp stings us, we are advised to rub soap over the sting? You must have observed that of *haldi* (turmeric) stains on our clothes changes in colour when we rub soap over it. A person suffering from acidity eats baking soda for relief. Lemon juice or tamarind is used to clean tarnished copper vessels. Think: Where can we see the effects of acids and bases in our daily lives?

Do you recall reading about indicators? Indicators help us classify substances into acidic, basic or neutral groups. For example, litmus is an indicator that changes colour in the presence of acids and alkalis. You also know that the reaction between an acid and a base produces a salt and water.

### 2.1 Acids and Bases are everywhere around us

Let us try to find out which acids and bases are present in the many substances we use in our daily life.

Table-1: Acids and Bases from Natural Sources

S.No.	Natural Source	Acid	S.No.	Natural Source	Base
1.	Tamarind	Tartaric acid	5.	Lime	Calcium hydroxide
2.	Apple	Malic acid	6.	Baking Soda	Sodium hydrogencarbonate
3.	Vinegar	Acetic acid	7.	Antacid	Magnesium hydroxide
4.	Tomato	Oxalic acid	8.	Washing soda	Sodium carbonate

In addition to the acids given in Table-1, we also have mineral/non-carbonic acids such as nitric acid ( $\text{HNO}_3$ ), sulphuric acid ( $\text{H}_2\text{SO}_4$ ) and hydrochloric acid ( $\text{HCl}$ ). Similarly, there are many other alkalis such as ammonium hydroxide ( $\text{NH}_4\text{OH}$ ), sodium hydroxide ( $\text{NaOH}$ ), potassium hydroxide ( $\text{KOH}$ ) etc.

#### Alkalis and Bases

Not all bases are soluble in water. Those bases that can be dissolved in water are called alkalis.

# Guided Reading And Study Workbook Chapter 2 acids Bases

**AW Rasmussen**



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