

Name:

Date:

Heats of Reaction

In this experiment you will determine and compare the quantity of heat energy released in three exothermic chemical reactions. Go to <http://chemcollective.org/vlab/138>

Reaction 1: Solid sodium hydroxide dissolved in water to form an aqueous solution of ions



Reaction 2: Solid sodium hydroxide reacts with an aqueous solution of hydrogen chloride to form water and an aqueous solution of sodium chloride.



Reaction 3: An aqueous solution of sodium hydroxide reacts with an aqueous solution of hydrogen chloride to form water and an aqueous solution of sodium chloride.



In order to accurately measure the heat released in each reaction, we will be using a calorimeter (a styrofoam cup). The change in temperature that occurs for each reaction will be used to calculate the energy released in kilojoules per mole of sodium hydroxide used. We can assume for our calculations that any heat transferred to the styrofoam and surrounding air will be negligible. We can also assume that the specific heat of water is $4.18 \text{ J/g}^\circ\text{C}$.

Specific heat can be calculated using the formula

$$q = mc\Delta T$$
$$q = -\Delta H$$

Procedure

Reaction 1

- In the glassware menu, take out a 50 ml graduated cylinder and a styrofoam cup. From the tools menu, take out the scale. From the chemical stockroom, move the distilled water and solid NaOH onto the workbench.
- Transfer 50 mL of distilled water to the styrofoam cup. To do this, drag the carboy of water onto the graduated cylinder. (Before you release the mouse button, the cursor will show a plus sign and a transfer text bar will appear. Enter 50.0 mL and click on pour. You will notice that the graduated cylinder reads 50.0 mL.
- Weigh about 1 g of solid sodium hydroxide pellets, $\text{NaOH}_{(s)}$, directly into the styrofoam cup and record its mass to the nearest 0.01 g. To do this, place the styrofoam cup on the balance so it registers a mass, then click the "Tare" button. Drag the NaOH button onto the styrofoam cup. (When you release the mouse, the bottle will be tipped to show that it is in pour mode.) Next type "1.00" g into the transfer bar and then click pour. Note that the balance now reads the mass of the transferred NaOH. You may now take the cup off of the scale and remove the scale from the workbench.

Heat Reaction Lab Answers

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