

VERY SHORT ANSWER TYPE QUESTIONS

2. Define mole fraction. 2
3. A first order reaction has rate constant of $7.39 \times 10^{-5} \text{ s}^{-1}$. Find the half life of this reaction. 2
4. Calculate the magnetic moment of Fe^{2+} (At. No. 26). 2
5. Write the IUPAC name of $[\text{Cr}(\text{NH}_3)_3\text{BrCl}_2]$. 2
6. How is ethyl iodide prepared from ethanol? 2
7. Discuss Reimer-Tiemann reaction. 2
8. Write Cannizzaro's reaction. 2
9. What happens when nitrobenzene is reduced with LiAlH_4 ? 2

SECTION-C

[SHORT ANSWER TYPE QUESTIONS]

10. Explain depression in freezing point. Show that it is a colligative property. 3
11. Standard electrode potentials are given as under :
 $\text{Ti}^{4+} + e^- \longrightarrow \text{Ti}^{3+} \quad E^\circ = + 0.01 \text{ V}$
 $\text{Fe}^{3+} + e^- \longrightarrow \text{Fe}^{2+} \quad E^\circ = + 0.77 \text{ V}$
Tell whether Ti^{4+} ion may be used to oxidise Fe^{2+} to Fe^{3+} . 3
12. Calculate the $t_{1/2}$ for a first order reaction, if the rate constant is 200 sec^{-1} . 3

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