

Name _____

answer key

Date _____

Stoichiometry: Mass - Mole Worksheet

Based on the following chemical equation, please answer the following questions!



Calculate the amount of moles of NaCl needed to produce 75 g of FeCl₂.

$$\frac{9 \text{ FeCl}_2}{1 \text{ mole FeCl}_2} \times \frac{1 \text{ mole FeCl}_2}{9 \text{ FeCl}_2} = \frac{\text{mole NaCl}}{\text{mole FeCl}_2}$$

Calculate the amount of Fe₂(SO₄)₃ in grams produce 215 g Na₂SO₄.

$$\frac{9 \text{ Na}_2\text{SO}_4}{1 \text{ mole Na}_2\text{SO}_4} \times \frac{1 \text{ mole Na}_2\text{SO}_4}{9 \text{ Na}_2\text{SO}_4} = \frac{\text{mole Fe}_2(\text{SO}_4)_3}{\text{mole Na}_2\text{SO}_4} \times \frac{9 \text{ Fe}_2(\text{SO}_4)_3}{1 \text{ mole Fe}_2(\text{SO}_4)_3}$$

Calculate the amount in grams of FeCl₂ produced with 95g of Fe (SO₄)₂.

$$\frac{9 \text{ Fe}_2(\text{SO}_4)_3}{1 \text{ mole Fe}_2(\text{SO}_4)_3} \times \frac{1 \text{ mole Fe}_2(\text{SO}_4)_3}{9 \text{ Fe}_2(\text{SO}_4)_3} = \frac{\text{mole FeCl}_2}{\text{mole Fe}(\text{SO}_4)_2} \times \frac{9 \text{ FeCl}_2}{1 \text{ mole FeCl}_2}$$

Calculate the amount of moles of Fe₂(SO₄)₃ needed to react with 123.5 g of NaCl

$$\frac{9 \text{ NaCl}}{1 \text{ mole NaCl}} \times \frac{1 \text{ mole NaCl}}{9 \text{ NaCl}} = \frac{\text{mole Fe}(\text{SO}_4)_2}{\text{mole NaCl}}$$

Based on the following chemical equation, please answer the following questions!



$$\frac{9 \text{ Ca(HCO}_3)_2}{1 \text{ mole Ca(HCO}_3)_2} \times \frac{1 \text{ mole Ca(HCO}_3)_2}{9 \text{ Ca(HCO}_3)_2} = \frac{\text{mole CaCO}_3}{\text{mole Ca(HCO}_3)_2}$$

Mole And Stoichiometry Practice Answer Key

Puja Mehta



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Table of Contents Mole And Stoichiometry Practice Answer Key

1. Understanding the eBook Mole And Stoichiometry Practice Answer Key
 - The Rise of Digital Reading Mole And Stoichiometry Practice Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Mole And Stoichiometry Practice Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mole And Stoichiometry Practice Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mole And Stoichiometry Practice Answer Key
 - Personalized Recommendations
 - Mole And Stoichiometry Practice Answer Key User Reviews and Ratings
 - Mole And Stoichiometry Practice Answer Key and Bestseller Lists
5. Accessing Mole And Stoichiometry Practice Answer Key Free and Paid eBooks

- Mole And Stoichiometry Practice Answer Key Public Domain eBooks
 - Mole And Stoichiometry Practice Answer Key eBook Subscription Services
 - Mole And Stoichiometry Practice Answer Key Budget-Friendly Options
6. Navigating Mole And Stoichiometry Practice Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Mole And Stoichiometry Practice Answer Key Compatibility with Devices
 - Mole And Stoichiometry Practice Answer Key Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mole And Stoichiometry Practice Answer Key
 - Highlighting and Note-Taking Mole And Stoichiometry Practice Answer Key
 - Interactive Elements Mole And Stoichiometry Practice Answer Key
 8. Staying Engaged with Mole And Stoichiometry Practice Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mole And Stoichiometry Practice Answer Key
 9. Balancing eBooks and Physical Books Mole And Stoichiometry Practice Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mole And Stoichiometry Practice Answer Key
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Mole And Stoichiometry Practice Answer Key
 - Setting Reading Goals Mole And Stoichiometry Practice Answer Key
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Mole And Stoichiometry Practice Answer Key
 - Fact-Checking eBook Content of Mole And Stoichiometry Practice Answer Key
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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