

Molarity Worksheet

Name: _____

1. What does molarity mean?

$$\frac{\text{Number of moles of solute}}{\text{1 liter solution}}$$

2. What is the molarity of a solution that contains 4.53 moles of lithium nitrate in 2.85 liters of solution?

$$\frac{4.53 \text{ mol LiNO}_3}{2.85 \text{ L soln}} = 1.59 \text{ M LiNO}_3$$

3. What is the molarity of a solution that contains 0.00372 moles hydrochloric acid in 2.39×10^{-2} liters of solution?

$$\frac{0.00372 \text{ mol HCl}}{2.39 \times 10^{-2} \text{ L soln}} = 0.156 \text{ M HCl}$$

4. A flask contains 85.5 g $\text{C}_{12}\text{H}_{22}\text{O}_{11}$ (sucrose) in 1.00 liters of solution. What is the molarity?

$$\frac{85.5 \text{ g sucrose} \times 1 \text{ mol sucrose}}{1.00 \text{ L soln} \quad 342.34 \text{ g sucrose}} = 0.250 \text{ M sucrose}$$

5. A beaker contains 214.2 grams osmium (III) fluoride in 0.0673 liters of solution. What is the molarity?

$$\frac{214.2 \text{ g OsF}_3 \times 1 \text{ mol OsF}_3}{0.0673 \text{ L soln} \quad 247.23 \text{ g OsF}_3} = 12.9 \text{ M OsF}_3$$

6. Calculate the molarity if a flask contains 1.54 moles potassium sulfate in 125 ml of solution.

$$\frac{1.54 \text{ mol K}_2\text{SO}_4}{0.125 \text{ L soln}} = 12.3 \text{ M K}_2\text{SO}_4$$

7. A beaker contains 36.45 grams ammonium chlorite in 2.36 liters of solution - calculate the molarity.

$$\frac{36.45 \text{ g NH}_4\text{ClO}_2 \times 1 \text{ mol NH}_4\text{ClO}_2}{2.36 \text{ L soln} \quad 85.50 \text{ g NH}_4\text{ClO}_2} = 0.181 \text{ M NH}_4\text{ClO}_2$$

8. What is the molarity of a solution that contains 14.92 grams magnesium oxalate in 3.65 ml of solution?

$$\frac{14.92 \text{ g MgC}_2\text{O}_4 \times 1 \text{ mol MgC}_2\text{O}_4}{0.00365 \text{ L soln} \quad 112.32 \text{ g MgC}_2\text{O}_4} = 36.4 \text{ mol MgC}_2\text{O}_4$$

Molarity Pre Lab Worksheet

Daniel F McAuley



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use Microsoft Excel to prepare and present data in graphical format and how to calculate their data into meaningful numbers that can be compared across experiments The second module requires that the students employ UV mutagenesis to generate hyper pigmented mutants of *S marcescens* for further characterization Students use experimental data and protocols learned in the first module to help them develop their own hypotheses experimental protocols and to analyze their own data Before each lab students are required to answer questions designed to probe their understanding of required pre laboratory reading materials Questions also guide the students through the development of hypotheses and predictions Following each laboratory students then answer a series of post laboratory questions to guide them through the presentation and analysis of their data and how to place their data into the context of primary literature Students are also asked to review their initial hypotheses and predictions to determine if their conclusions are supportive A formal laboratory report is also to be completed after each module in a format similar to that of primary scientific literature The Fundamentals of Scientific Research An Introductory Laboratory Manual is an invaluable resource to undergraduates majoring in the life sciences

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