

Molarity Practice problems

$$M = \frac{\text{mol}}{L}$$



Key

1. How many moles of CaCl_2 would be dissolved in 1.0 L of a 0.1 M solution of CaCl_2 ?

$$\text{mol} = M \times L = 0.1 \frac{\text{mol}}{L} \times 1 L = \boxed{0.1 \text{ mol } \text{CaCl}_2}$$

2. What is the concentration of 25 mL of solution that has 0.05 moles of solute dissolved in it?

$$M = \frac{\text{mol}}{L} = \frac{0.05 \text{ mol}}{0.025 L} = \boxed{2 M}$$

3. How many liters of water need to be added to 0.5 moles of CaCl_2 to make a 0.20 M solution of CaCl_2 ?

$$L = \frac{\text{mol}}{M} = \frac{0.5 \text{ mol}}{0.20 \frac{\text{mol}}{L}} = \boxed{2.5 L}$$

4. How many grams of potassium carbonate are needed to make 200 mL of a 2.5 M solution?

$$\text{mol} = M \times L \quad \text{K}_2\text{CO}_3$$

$$\text{mol} = 2.5 M \times .200 L = 0.50 \text{ mol}$$

$$\frac{0.50 \text{ mol } \text{K}_2\text{CO}_3 \mid 138.21 \text{ g}}{1 \text{ mol}} = \boxed{69.1 \text{ g } \text{K}_2\text{CO}_3}$$

5. How many grams of Lithium Bromide would be needed to make 1.43 L of 4 M Lithium Bromide solution?

$$\text{mol} = M \times L \quad \text{LiBr}$$

$$= 4 M \times 1.43 L$$

$$= 5.72 \text{ mol LiBr}$$

$$\frac{5.72 \text{ mol LiBr} \mid 86.841 \text{ g}}{1 \text{ mol}} = \boxed{496.7 \text{ g LiBr}}$$

6. What is the concentration of a 450 mL solution that contains 200 grams of iron (II) chloride? FeCl_2

$$M = \frac{\text{mol}}{L}$$

$$\frac{200 \text{ g } \text{FeCl}_2 \mid 1 \text{ mol}}{126.75 \text{ g}} = 1.578 \text{ mol } \text{FeCl}_2$$

$$\frac{1.578 \text{ mol}}{.450 L} = \boxed{3.5 M}$$

7. How many grams of ammonium sulfate are needed to make a 0.25 L solution at a concentration of 6 M?

$$\text{mol} = M \times L \quad (\text{NH}_4)_2\text{SO}_4$$

$$= 6 M \times .25 L$$

$$= 1.5 \text{ mol } (\text{NH}_4)_2\text{SO}_4$$

$$\frac{1.5 \text{ mol } (\text{NH}_4)_2\text{SO}_4 \mid 132.138 \text{ g}}{1 \text{ mol}} = \boxed{198.2 \text{ g } (\text{NH}_4)_2\text{SO}_4}$$

8. What is the concentration of a solution that has a volume of 2.5 L and contains 660 grams of calcium phosphate?

$$M = \frac{\text{mol}}{L} \quad \text{Ca}_3(\text{PO}_4)_2$$

$$\frac{660 \text{ g } \text{Ca}_3(\text{PO}_4)_2 \mid 1 \text{ mol}}{310.18 \text{ g}} = 2.127 \text{ mol}$$

$$\frac{2.127 \text{ mol}}{2.5 L} = \boxed{0.85 M}$$

9. How many grams of copper (II) fluoride are needed to make 6.7 liters of a 1.2 M solution?

$$\text{mol} = M \times L \quad \text{CuF}_2$$

$$= 1.2 M \times 6.7 L$$

$$= 8.04 \text{ mol } \text{CuF}_2$$

$$\frac{8.04 \text{ mol } \text{CuF}_2 \mid 101.55 \text{ g}}{1 \text{ mol}} = \boxed{816.5 \text{ g } \text{CuF}_2}$$

Molarity Pogil Lab Answer Key

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Molarity Pogil Lab Answer Key:

Unveiling the Magic of Words: A Review of "**Molarity Pogil Lab Answer Key**"

In a global defined by information and interconnectivity, the enchanting power of words has acquired unparalleled significance. Their power to kindle emotions, provoke contemplation, and ignite transformative change is really awe-inspiring. Enter the realm of "**Molarity Pogil Lab Answer Key**," a mesmerizing literary masterpiece penned by a distinguished author, guiding readers on a profound journey to unravel the secrets and potential hidden within every word. In this critique, we shall delve to the book is central themes, examine its distinctive writing style, and assess its profound impact on the souls of its readers.

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