

1/5/16

## Mole Ratio Worksheet

Key

1) Given this equation:  $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$ , write the following molar ratios:

- a)  $\text{N}_2 / \text{H}_2$  1:3  
 b)  $\text{N}_2 / \text{NH}_3$  1:2  
 c)  $\text{H}_2 / \text{NH}_3$  3:2

2) Given the following equation:  $8\text{H}_2 + \text{S}_8 \rightarrow 8\text{H}_2\text{S}$ , write the following molar ratios:

- a)  $\text{H}_2 / \text{H}_2\text{S}$  8:8  $\rightarrow$  1:1 (reduced)  
 b)  $\text{H}_2 / \text{S}_8$  8:1  
 c)  $\text{H}_2\text{S} / \text{S}_8$  8:1

3) Answer the following questions for this equation:  $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$

- a) What is the  $\text{H}_2 / \text{H}_2\text{O}$  molar ratio? 2:2  $\rightarrow$  1:1 (reduced)  
 b) Suppose you had 20 moles of  $\text{H}_2$  on hand and plenty of  $\text{O}_2$ , how many moles of  $\text{H}_2\text{O}$  could you make?  

$$\frac{20 \text{ mole H}_2}{2 \text{ H}_2} \times \frac{2 \text{ H}_2\text{O}}{2 \text{ H}_2} = \boxed{20 \text{ mole H}_2\text{O}}$$
  
 c) What is the  $\text{O}_2 / \text{H}_2\text{O}$  molar ratio?  
 1:2  
 d) Suppose you had 20 moles of  $\text{O}_2$  and enough  $\text{H}_2$ , how many moles of  $\text{H}_2\text{O}$  could you make?  

$$\frac{20 \text{ mol O}_2}{1 \text{ O}_2} \times \frac{2 \text{ H}_2\text{O}}{2 \text{ H}_2} = \boxed{40 \text{ mol H}_2\text{O}}$$

4) Use this equation:  $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$ , for the following problems

- a) If you used 1 mole of  $\text{N}_2$ , how many moles of  $\text{NH}_3$  could be produced?  $\boxed{2}$   
 b) If 10 moles of  $\text{NH}_3$  were produced, how many moles of  $\text{N}_2$  would be required?  

$$\frac{10 \text{ mol NH}_3}{2 \text{ NH}_3} \times \frac{1 \text{ N}_2}{1 \text{ N}_2} = \boxed{5 \text{ mol N}_2}$$
  
 c) If 3.00 moles of  $\text{H}_2$  were used, how many moles of  $\text{NH}_3$  would be made?  

$$\frac{3 \text{ mol H}_2}{3 \text{ H}_2} \times \frac{2 \text{ NH}_3}{2 \text{ NH}_3} = \boxed{2 \text{ mol NH}_3}$$
  
 d) If 0.600 moles of  $\text{NH}_3$  were produced, how many moles of  $\text{H}_2$  are required?  

$$\frac{0.6 \text{ mol NH}_3}{2 \text{ NH}_3} \times \frac{3 \text{ H}_2}{1 \text{ N}_2} = \boxed{0.9 \text{ mol H}_2}$$

# Finding Mass Using Mole Ratio Answers

**Deborah Wiegand**



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