

UNIT TEST 1

(2025-2026)

GRADE XIIA,B

DATE: 14.02.25

CHEMISTRY

TIME: Hr

MARKS: 25

I. Answer the following questions :-

8X1=8

1. The value of Henry's constant K_H is:

- a) greater for gases with higher solubility
- b) greater for gases with lower solubility.
- c) constant for all gases.
- d) not related to the solubility of gases.

2. Increasing the temperature of an aqueous solution will cause:

- a) decrease in molality
- b) decrease in molarity
- c) decrease in mole fraction
- d) decrease in % (w/w)

3. If liquid A and B form ideal solution:

- a) The enthalpy of mixing is zero
- b) The entropy of mixing is zero
- c) The free energy of mixing is zero
- d) free energy as well as entropy of mixing are zero

4. Which of the following condition is not satisfied by an ideal solution?

- a) $\Delta H_{\text{mixing}} = 0$
- b) $\Delta V_{\text{mixing}} = 0$
- c) Raoult's Law is obeyed
- d) Formation of an azeotropic mixture

5. Which of the following pair will make an ideal solution?

- a) Chlorobenzene + Chloroethane
- b) Benzene + Toluene
- c) Acetone + Chloroform
- d) water + HCl

Given below are two statements labelled as Assertion and Reason

Select the most appropriate answer from the options given below:

- a) Both A and R are true and R is the correct explanation of A
- b) Both A and R are true but R is not the correct explanation of A
- c) A is true but R is false
- d) A is false but R is true

6. Assertion : Molarity of a solution in liquid state changes with temperature.

Reason : The volume of a solution changes with change in temperature.

7. Assertion : When methyl alcohol is added to water, boiling point of water increases.

Reason : When a volatile solute is added to a volatile solvent elevation in boiling point is observed.

8. Assertion: Azeotropic mixture are formed only by non ideal solutions and they may boiling points either greater than both the components or less than both the components.

Reason : The composition of the vapour phase is same as that of liquid phase of an azeotropic mixture.

II. Answer the following question:-

1X2=2

9. Write down four important points of differences between an ideal and a non ideal solution.

OR

What do you mean by Raoult's law? What are the limitations of Raoult's law?

III. Answer the following questions:-

2X3=6

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