

Limiting Reactant Practice

5 pages + 5-page Answer KEY + Periodic Table



LIMITING REACTANTS

1. The reaction below produces ammonia for use in fertilizers.
 $N_2 + 3H_2 \rightarrow 2NH_3$
If you start with 28 grams of N_2 and 12 grams of H_2 , determine the limiting reactant.

2. Calculate the grams of ammonia produced.

3. How many grams of the excess reactant is produced in a left over?
* Multiply the mass of the limiting reactant by the mole ratio of the excess reactant.

4. Calculate moles of product * mole of reactant.

2. The reaction below produces carbon dioxide and water for a gas.
 $C_2H_4 + 3O_2 \rightarrow 2CO_2 + 2H_2O$
If you have 10 grams of C_2H_4 and 60 grams of O_2 , determine the limiting reactant.

3. How many grams of water produced are produced?



4. Magnesium chloride is used to create water for drinking tanks. The reaction is:
 $Mg + 2HCl \rightarrow MgCl_2 + H_2$
If you start with 24 grams of magnesium and 73 grams of hydrochloric acid, determine the limiting reactant.

5. Calculate the grams of magnesium chloride produced.

6. How many grams of the excess reactant is produced in a left over?

7. The reaction below is used to create water for drinking tanks. The reaction is:
 $2H_2 + O_2 \rightarrow 2H_2O$
If you have 10 grams of H_2 and 60 grams of O_2 , determine the limiting reactant.

8. How many grams of water produced are produced?

9. The reaction below is used to create water for drinking tanks. The reaction is:
 $2H_2 + O_2 \rightarrow 2H_2O$
If you have 10 grams of H_2 and 60 grams of O_2 , determine the limiting reactant.

10. Calculate the grams of water produced.

11. How many grams of the excess reactant is produced in a left over?

12. Calculate the grams of water produced.

13. How many grams of the excess reactant is produced in a left over?

14. Calculate the grams of water produced.

15. How many grams of the excess reactant is produced in a left over?

ANSWER KEY INCLUDED

Limiting Reactant Answer Key

Rosina Ehmman



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Table of Contents Limiting Reactant Answer Key

1. Understanding the eBook Limiting Reactant Answer Key
 - The Rise of Digital Reading Limiting Reactant Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Limiting Reactant Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Limiting Reactant Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Limiting Reactant Answer Key
 - Personalized Recommendations
 - Limiting Reactant Answer Key User Reviews and Ratings
 - Limiting Reactant Answer Key and Bestseller Lists
5. Accessing Limiting Reactant Answer Key Free and Paid eBooks
 - Limiting Reactant Answer Key Public Domain eBooks
 - Limiting Reactant Answer Key eBook Subscription Services
 - Limiting Reactant Answer Key Budget-Friendly Options
6. Navigating Limiting Reactant Answer Key eBook Formats

- ePub, PDF, MOBI, and More
 - Limiting Reactant Answer Key Compatibility with Devices
 - Limiting Reactant Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Limiting Reactant Answer Key
 - Highlighting and Note-Taking Limiting Reactant Answer Key
 - Interactive Elements Limiting Reactant Answer Key
8. Staying Engaged with Limiting Reactant Answer Key
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Limiting Reactant Answer Key
9. Balancing eBooks and Physical Books Limiting Reactant Answer Key
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Limiting Reactant Answer Key
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Limiting Reactant Answer Key
- Setting Reading Goals Limiting Reactant Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Limiting Reactant Answer Key
- Fact-Checking eBook Content of Limiting Reactant Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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