

## Electron Configuration Practice Worksheet **KEY**

In the space below, write the unabbreviated electron configurations of the following elements:

- 1) sodium  $1s^2 2s^2 2p^6 3s^1$
- 2) iron  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^6$
- 3) bromine  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^5$
- 4) barium  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2$
- 5) neptunium  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^6 7s^2 5f^4$

In the space below, write the abbreviated electron configurations of the following elements:

- 6) cobalt  $[\text{Ar}] 4s^2 3d^7$
- 7) silver  $[\text{Kr}] 5s^2 4d^9$
- 8) tellurium  $[\text{Kr}] 5s^2 4d^{10} 5p^4$
- 9) radium  $[\text{Rn}] 7s^2$
- 10) lawrencium  $[\text{Rn}] 7s^2 5f^{14} 6d^1$

Determine what elements are denoted by the following electron configurations:

- 11)  $1s^2 2s^2 2p^6 3s^2 3p^4$  **Sulfur**
- 12)  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1$  **Rubidium**
- 13)  $[\text{Kr}] 5s^2 4d^{10} 5p^3$  **Antimony**
- 14)  $[\text{Xe}] 6s^2 4f^{14} 5d^8$  **Osmium**
- 15)  $[\text{Rn}] 7s^2 5f^{11}$  **Einsteinium**

Determine which of the following electron configurations are not valid:

- 16)  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 4d^{10} 4p^5$  **should be  $3d^{10}$**
- 17)  $1s^2 2s^2 2p^6 3s^2 3d^5$   **$3s^2$  is impossible & "p" instead of "d"**
- 18)  $[\text{Ra}] 7s^2 5f^8$  **Ra is not a noble gas**
- 19)  $[\text{Kr}] 5s^2 4d^{10} 5p^5$  **Valid**
- 20)  $[\text{Xe}]$  **Element can't be its own electron configuration.**

# Electron Configuration Worksheet Answer Key

**Dale Rich**



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