

# Gas Laws Study Guide

## Test Format:

- Multiple choice and problems

## Items that will be provided for the test (read important notes at the bottom):

- Gas law equations\*
  - The value of R.
  - Pressure conversion factors
- \* I will provide the equations on the test. You are responsible for knowing which laws the equations describe (ex: I will give you  $P_1 \times V_1 = P_2 \times V_2$  and you must know that this is Boyle's Law).

- What is the name of the model we currently use to explain the behavior of gases? **kinetic molecular theory (KMT)**
- Define the kinetic molecular theory (KMT). **Gases are in constant motion with elastic collisions**
- The KMT makes many assumptions about the Size, motion, and energy of gas particles.
- According to the KMT, do gas particles experience attractive or repulsive forces? Explain why or why not. **Neither- they are too far apart**
- Completely describe the motion of gas particles according to the KMT. **Constant, Random, Straight paths, different velocity**
- Define elastic collision. **No kinetic energy is lost- just transferred between particles**
- What determines the kinetic energy of a particle? **Mass and velocity**
- In a sample of a single gas, do all particles have the same kinetic energy? Explain why or why not. **No - they have a different velocity**
- Define temperature. **Measurement of kinetic energy**
- Explain why gases have a low density. **Lots of space between particles (large volume)**
- Why are gases compressible? **Lots of empty space between particles**
- Why do gases expand to fill a container? **No attractive or repulsive forces**
- Why can gas particles flow easily past each other? **No attractive forces**
- Define diffusion. **Movement of one material through another**
- Define effusion. **Gas escapes through tiny openings**
- Describe how diffusion depends on the mass of the particles. **Diffusion is dependent on mass, so less mass more diffusion**
- Use the KMT to explain why a balloon expands as you blow air into it. **Particles are in constant motion and collide with their surroundings expanding the balloon**

## Boyle's Law: Answer the following questions about Boyle's Law.

- Describe Boyle's Law in your own words. **As pressure increases, volume decreases**
- What is the formula?  **$P_1V_1 = P_2V_2$**
- What are the variables? What are the possible units used in the formula? **Pressure and volume, atm, kPa, mL, L**
- What is constant? **temperature**



- Give an example to help explain the law (you should have some type of visual of your example).
- What would a graph of pressure vs. volume of a gas look like according to Boyle's Law?

## Charles' Law: Answer the following questions about Charles' Law.

- Describe Charles' Law in your own words. **Increase temperature, increase volume**
- What is the formula?  **$V = kT$**
- What are the variables? What are the possible units used in the formula? **Temperature and volume (kelvin, mL, L)**
- What is constant? **Pressure**
- What is very important to remember about the temperature variable when using this formula? **It must be in kelvin**
- Give an example to help explain Charles's Law (you should have some type of visual of your example).



- What would a graph of temperature vs. volume of a gas look like according to Charles's Law?

## Gay-Lussac's Law: Answer the following questions about Gay-Lussac's Law.

- Describe Gay-Lussac's Law in your own words. **As pressure increases, temperature increases**
- What is the formula?  **$P = kT$**
- What are the variables? What are the possible units used in the formula? **Pressure and temperature (kelvin, atm, kPa, torr)**
- What is constant? **volume**
- What is very important to remember about the temperature variable when using this formula? **It must be in kelvin**
- Give an example to help explain Gay-Lussac's Law (you should have some type of visual of your example).

# Gas Study Guide

**Michael J. Moran, Howard N.  
Shapiro, Daisy D. Boettner, Margaret  
B. Bailey**

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Laboratory Directions and Study Questions in Inorganic Chemistry Alexander Silverman, 1919

*Study Guide and Student Solutions Manual to Accompany Physics for Scientists and Engineers, by Serway* Raymond A. Serway, Steven Van Wyk, 1996

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**Study Guide and Exercises for Welding Processes and Power Sources** Edward R. Pierre, 1985

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The Gas Record , 1921

**Petroleum Register** , 1922

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**Reader's Guide to Periodical Literature Supplement** , 1916 These vols contain the same material as the early vols of Social sciences humanities index

Journal of Gas Lighting and Water Supply , 1909

Respiratory Care , 1986

*A Manual of the High-speed Steam Engine* Harry Keay Pratt, 1914

**Gas Journal** , 1926

Fundamentals of Engineering Thermodynamics Michael J. Moran, Howard N. Shapiro, Daisie D. Boettner, Margaret B. Bailey, 2010-12-07 This leading text in the field maintains its engaging readable style while presenting a broader range of applications that motivate engineers to learn the core thermodynamics concepts Two new coauthors help update the material and integrate engaging new problems Throughout the chapters they focus on the relevance of thermodynamics to modern engineering problems Many relevant engineering based situations are also presented to help engineers model and solve these problems

The American Gas Light Journal , 1905

Country Guide and Nor'west Farmer , 1911

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