

Name:

Date:

Heats of Reaction

In this experiment you will determine and compare the quantity of heat energy released in three exothermic chemical reactions. Go to <http://chemcollective.org/vlab/138>

Reaction 1: Solid sodium hydroxide dissolved in water to form an aqueous solution of ions



Reaction 2: Solid sodium hydroxide reacts with an aqueous solution of hydrogen chloride to form water and an aqueous solution of sodium chloride.



Reaction 3: An aqueous solution of sodium hydroxide reacts with an aqueous solution of hydrogen chloride to form water and an aqueous solution of sodium chloride.



In order to accurately measure the heat released in each reaction, we will be using a calorimeter (a styrofoam cup). The change in temperature that occurs for each reaction will be used to calculate the energy released in kilojoules per mole of sodium hydroxide used. We can assume for our calculations that any heat transferred to the styrofoam and surrounding air will be negligible. We can also assume that the specific heat of water is $4.18 \text{ J/g}^\circ\text{C}$.

Specific heat can be calculated using the formula

$$q = mc\Delta T$$
$$q = -\Delta H$$

Procedure

Reaction 1

- In the glassware menu, take out a 50 ml graduated cylinder and a styrofoam cup. From the tools menu, take out the scale. From the chemical stockroom, move the distilled water and solid NaOH onto the workbench.
- Transfer 50 mL of distilled water to the styrofoam cup. To do this, drag the carboy of water onto the graduated cylinder. (Before you release the mouse button, the cursor will show a plus sign and a transfer text bar will appear. Enter 50.0 mL and click on pour. You will notice that the graduated cylinder reads 50.0 mL.
- Weigh about 1 g of solid sodium hydroxide pellets, $\text{NaOH}_{(s)}$, directly into the styrofoam cup and record its mass to the nearest 0.01 g. To do this, place the styrofoam cup on the balance so it registers a mass, then click the "Tare" button. Drag the NaOH button onto the styrofoam cup. (When you release the mouse, the bottle will be tipped to show that it is in pour mode.) Next type "1.00" g into the transfer bar and then click pour. Note that the balance now reads the mass of the transferred NaOH. You may now take the cup off of the scale and remove the scale from the workbench.

Heat Reaction Lab Answers

Lingjun Ying



Heat Reaction Lab Answers:

The book delves into Heat Reaction Lab Answers. Heat Reaction Lab Answers is a vital topic that needs to be grasped by everyone, ranging from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Heat Reaction Lab Answers, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:
 - Chapter 1: Introduction to Heat Reaction Lab Answers
 - Chapter 2: Essential Elements of Heat Reaction Lab Answers
 - Chapter 3: Heat Reaction Lab Answers in Everyday Life
 - Chapter 4: Heat Reaction Lab Answers in Specific Contexts
 - Chapter 5: Conclusion
2. In chapter 1, the author will provide an overview of Heat Reaction Lab Answers. The first chapter will explore what Heat Reaction Lab Answers is, why Heat Reaction Lab Answers is vital, and how to effectively learn about Heat Reaction Lab Answers.
3. In chapter 2, the author will delve into the foundational concepts of Heat Reaction Lab Answers. This chapter will elucidate the essential principles that must be understood to grasp Heat Reaction Lab Answers in its entirety.
4. In chapter 3, this book will examine the practical applications of Heat Reaction Lab Answers in daily life. This chapter will showcase real-world examples of how Heat Reaction Lab Answers can be effectively utilized in everyday scenarios.
5. In chapter 4, this book will scrutinize the relevance of Heat Reaction Lab Answers in specific contexts. The fourth chapter will explore how Heat Reaction Lab Answers is applied in specialized fields, such as education, business, and technology.
6. In chapter 5, this book will draw a conclusion about Heat Reaction Lab Answers. The final chapter will summarize the key points that have been discussed throughout the book.

The book is crafted in an easy-to-understand language and is complemented by engaging illustrations. This book is highly recommended for anyone seeking to gain a comprehensive understanding of Heat Reaction Lab Answers.

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