

Exercise-6(A)

1. What is heat? Write its S.I unit.

Solution:

Heat is the internal energy of molecules constituting the body. It flows from a hot body to a cold body. The S.I. unit of heat is joule (J).

2. Two bodies at different temperatures are placed in contact. State the direction in which heat will flow.

Solution:

Heat flows from the body having higher temperature to a body with low temperature.

3. Name the S.I. unit of heat. How is it related to the unit caloric?

Solution:

The S.I. unit of heat is joule (J).

Joule is related to unit caloric in the following way:

$1\text{J} = 0.24\text{cal}$ (approximately)

4. Define temperature and write its S.I unit.

Solution:

Temperature is a quantity which conveys the thermal state of a body (i.e., the degree of hotness or coolness of the body). It determines the direction of flow of heat when two bodies at different temperatures are placed in contact.

The S.I. unit of temperature is Kelvin (K)

5. Why does a piece of ice when touched with hand, appear cool? Explain.

Solution:

It appears cool because when ice is touched, heat passes from our hand to the ice.

6. Distinguish between heat and temperature.

Solution:

The differences are as follows:

Heat	Temperature
Form of energy obtained as a result of random motion of molecules in a substance	Quantity that indicates the thermal state of a body, determines direction of heat flow when two bodies are kept in contact having different temperatures
S.I. unit - Joule (J)	S.I. unit - Kelvin (K)
Measured using - Principle of calorimetry	Measured using - Thermometer
Amount of heat contained in a body depends on mass, temperature, material of the body	Temperature depends on the average kinetic energy of its molecules as a result of their random motion

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