

Specific Heat Capacity Practice Problems

Name:

1. When 3.0 kg of water is cooled from 80.0°C to 10.0°C, how much heat energy is lost?

$$c_{p,w} = Q / m \cdot \Delta T$$

$$Q = c_{p,w} (m) \Delta T$$

$$= (4186)(3)(10 - 80)$$

$$= -879,060 \text{ J}$$

2. How much heat is needed to raise a 0.30 kg piece of aluminum from 30.°C to 150°C?

$$Q = c_{p,a} (m_a) (\Delta T_a)$$

$$= 899 (.3) (150 - 30)$$

$$= +32,364 \text{ J}$$

3. Calculate the temperature change when:

- a) 10.0 kg of water loses 232 kJ of heat.

$$232 \text{ kJ} \cdot \frac{1000 \text{ J}}{1 \text{ kJ}} = 232,000 \text{ J}$$

$$-232,000 \text{ J} = (4186)(10.0 \text{ kg}) \Delta T$$

$$\frac{-232,000}{4186} = \Delta T$$

$$-55.6^\circ\text{C} = \Delta T$$

- b) 1.96 kJ of heat are added to 500. g of copper.

$$1.96 \text{ kJ} \cdot \frac{1000 \text{ J}}{1 \text{ kJ}} = 1960 \text{ J}$$

$$500 \text{ g} \cdot \frac{1 \text{ kg}}{1000 \text{ g}} = .5 \text{ kg}$$

$$1960 \text{ J} = (387)(.5 \text{ kg}) \Delta T$$

$$\frac{1960}{193.5} = \Delta T$$

$$10.1^\circ\text{C} = \Delta T$$

4. $2.52 \times 10^4 \text{ J}$ of heat are added to 2.0 kg of mercury to reach a final temperature of 130°C. What was the initial temperature of the mercury?

$$T_f = 130^\circ\text{C}$$

$$Q = 2.52 \times 10^4$$

$$m = 2.0 \text{ kg}$$

$$c_{p,m} = 138$$

$$2.52 \times 10^4 = (138)(2)(130 - T_i)$$

$$2.52 \times 10^4 = 276(130 - T_i)$$

$$91.3 = 130 - T_i$$

$$-38.7 = -T_i$$

$$38.7^\circ\text{C} = T_i$$

Heat Change Practice Problems

JE Gale



Heat Change Practice Problems:

Heat Change Practice Problems Book Review: Unveiling the Magic of Language

In an electronic digital era where connections and knowledge reign supreme, the enchanting power of language has become much more apparent than ever. Its capability to stir emotions, provoke thought, and instigate transformation is really remarkable. This extraordinary book, aptly titled "**Heat Change Practice Problems**," written by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound impact on our existence. Throughout this critique, we will delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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