

Student Exploration: Stoichiometry

Gizmo Warm-up

Just as a cook follows a recipe to decide how much of each ingredient to add, a chemist uses **stoichiometry** to determine the amounts of substances involved in chemical reactions. The Stoichiometry Gizmo™ allows you to try your hand at figuring out the amounts of resources and products that take part in a chemical reaction.

To begin, check that this equation is shown:



1. Look at the **coefficients** (such as the “3” in 3CO_2) in front of each substance in the equation. The coefficients tell you how many molecules or atoms take part in a chemical reaction. In the spaces below, list the number of each molecule or atom in the equation:

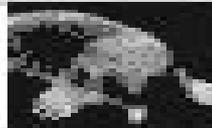
Fe_2O_3 _____ CO _____ Fe _____ CO_2 _____

2. In a **balanced equation**, the same number of each kind of atom is shown on each side of the equation. Calculate the number of iron (Fe), oxygen (O), and carbon atoms (C).

Reactants Iron: _____ Oxygen: _____ Carbon: _____

Products Iron: _____ Oxygen: _____ Carbon: _____

Based on these values, is the equation balanced? _____

Activity A: Moles	<p>Get the Gizmo ready: Check that the equation is still: $\text{Fe}_2\text{O}_3 + 3\text{CO} \rightarrow 2\text{Fe} + 3\text{CO}_2$ If not, click New equation until it reappears.</p>	
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Question: How do scientists find the molecular mass and molar mass of a substance?

Calculate: The **molecular mass** of a molecule is the sum of the masses of each atom in the molecule. The unit of molar mass is grams.

Iron's atomic mass is 55.8g, carbon's mass is 12.0g, and oxygen's mass is 16.0g.

Calculate the molecular mass of carbon monoxide (CO) by adding the atomic mass of carbon and the atomic mass of oxygen: _____

Calculate the molecular mass of carbon dioxide (CO₂): _____

Calculate the molecular mass of iron (II) oxide (Fe₂O₃): _____

Check your answers on the Gizmo by inspecting the middle row of tiles on the right side of the Gizmo. These tiles show the units “L/mol” on top and “g” below.

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