

Chapter 7: Ionic Compounds and Metals

7.1 Ion Formation p. 206

1. A **chemical bond** is the force that holds two atoms together.
2. Chemical bonds form by the attraction between the positive nucleus of one atom and the negative electrons of another atom.
3. Atom's try to form the **octet**—the stable arrangement of eight valence electrons in the outer energy level—by gaining or losing valence electrons.
4. A positively charged ion is called a **cation**.
5. An example is when sodium atom loses one valence electron to become a cation ion.
6. **Metals** are reactive because they lose valence electrons easily.
7. Transition metals commonly form 2+ or 3+ ions, but can form greater than 3+ ions.
8. Other relatively stable electron arrangements are referred to as **pseudo-noble gas configurations**.
9. An **anion** is a negatively charged ion.
10. An example would be chlorine gaining an electron to become a chlorine ion.
11. Nonmetal ions gain the number of electrons required to fill an octet.
12. Some nonmetals can gain or lose electrons to complete an octet.

7.2 Ionic Bonds and Ionic Compounds p.210

1. The electrostatic force that holds oppositely charged particles together in an ionic compound is called an **ionic bond**.
2. Compounds that contain ionic bonds are called **ionic compounds**.
3. Binary ionic compounds contain only two different elements—a metallic **cation** and a nonmetallic **anion**.
4. Formation of an ionic bond (table 7.4 p.211)
5. Positive and negative ions exist in a ratio determined by the number of electrons transferred from the metal atom to the non-metal atom.
6. The repeating pattern of particle packing in an ionic compound is called an **ionic crystal**.
7. The strong attractions among the positive and negative ions result in the formation of the crystal lattice.
8. A **crystal lattice** is the three-dimensional geometric arrangement of particles, and is responsible for the structure of many minerals.
9. Melting point, boiling point, and hardness depend on the strength of the attraction.
10. Melting and boiling points of some ionic compounds table 7.5 p.215.
11. In a **solid, ions are locked** into position and electrons **cannot flow freely**—solid ions are **poor conductors of electricity**.
12. **Liquid ions** or ions in aqueous solution have electrons that are **free to move**, so they **conduct electricity easily**.
13. An ion in aqueous solution that conducts electricity is an **electrolyte**.
14. Reactions that absorb energy are **endothermic**.
15. Reactions that release energy are **exothermic**.
16. The energy required to separate 1 mol of ions in an ionic compound is referred to as the **lattice energy**.
17. Lattice energy is **directly related** to the size of the ions that are bonded.

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