

Acids and Bases Worksheet

1. Acids and bases share some properties but are different in others. Use the samples at the side of the class to complete the chart. If both acids and bases share the property, place it in both columns.

Property	Acids	Bases
Conductivity in Aqueous Solution	Very	Very
Taste	Sour	Bitter
Feel	—	Slippery
Reaction with Metal	produces H_2 gas	—
Reaction with Carbonate	produces CO_2 gas	—
Colour on Litmus Paper	Red	Blue
Colour with Universal Indicator	Red	Purple
Colour with phenolphthalein	Colourless	pink

2. Indicate if each compound is an acid or base and name it.

	Acid/Base?	Name
HNO_3	<u>A</u>	<u>nitric acid</u>
$Mg(OH)_2$	<u>B</u>	<u>magnesium hydroxide</u>
H_3PO_4	<u>A</u>	<u>phosphoric acid</u>
H_2CO_3	<u>A</u>	<u>carbonic acid</u>
HF	<u>A</u>	<u>hydrofluoric acid</u>
$Fe(OH)_3$	<u>B</u>	<u>iron (III) hydroxide</u>

3. Write the formula for each of the following acids/bases:

hydrochloric acid	<u>HCl</u>	hydrochloric acid	<u>HCl</u>
aluminum hydroxide	<u>$Al(OH)_3$</u>	sodium hydroxide	<u>NaOH</u>
lithium hydroxide	<u>LiOH</u>	calcium hydroxide	<u>$Ca(OH)_2$</u>
sulfuric acid	<u>H_2SO_4</u>	sulfurous acid	<u>H_2SO_3</u>

Introduction To Acids And Bases Answer Key

Annelies Wilder-Smith



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web key points potential energy is energy that has the potential to become another form of energy an object's potential energy depends on its physical properties and position in a system potential energy comes in many forms such as gravitational potential energy due to an object's mass and position in a gravitational field

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web the energy of the reactants is shown on the energy diagram above on the right side of the energy diagram draw a horizontal line segment to indicate the energy of the products draw a vertical line segment to indicate Δh for the reaction and label it with the correct value see attached image

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