

# Gas Laws Study Guide

## Test Format:

- Multiple choice and problems

## Items that will be provided for the test (read important notes at the bottom):

- Gas law equations\*
  - The value of R.
  - Pressure conversion factors
- \* I will provide the equations on the test. You are responsible for knowing which laws the equations describe (ex: I will give you  $P_1 \times V_1 = P_2 \times V_2$  and you must know that this is Boyle's Law).

- What is the name of the model we currently use to explain the behavior of gases? **kinetic molecular theory (KMT)**
- Define the kinetic molecular theory (KMT). **Gases are in constant motion with elastic collisions**
- The KMT makes many assumptions about the Size, motion, and energy of gas particles.
- According to the KMT, do gas particles experience attractive or repulsive forces? Explain why or why not. **Neither- they are too far apart**
- Completely describe the motion of gas particles according to the KMT. **Constant, Random, Straight paths, different velocity**
- Define elastic collision. **No kinetic energy is lost- just transferred between particles**
- What determines the kinetic energy of a particle? **Mass and velocity**
- In a sample of a single gas, do all particles have the same kinetic energy? Explain why or why not. **No - they have a different velocity**
- Define temperature. **Measurement of kinetic energy**
- Explain why gases have a low density. **Lots of space between particles (large volume)**
- Why are gases compressible? **Lots of empty space between particles**
- Why do gases expand to fill a container? **No attractive or repulsive forces**
- Why can gas particles flow easily past each other? **No attractive forces**
- Define diffusion. **Movement of one material through another**
- Define effusion. **Gas escapes through tiny openings**
- Describe how diffusion depends on the mass of the particles. **Diffusion is dependent on mass, so less mass more diffusion**
- Use the KMT to explain why a balloon expands as you blow air into it. **Particles are in constant motion and collide with their surroundings expanding the balloon**

## Boyle's Law: Answer the following questions about Boyle's Law.

- Describe Boyle's Law in your own words. **As pressure increases, volume decreases**
- What is the formula?  **$P_1V_1 = P_2V_2$**
- What are the variables? What are the possible units used in the formula? **Pressure and volume, atm, kPa, mL, L**
- What is constant? **temperature**



- Give an example to help explain the law (you should have some type of visual of your example).
- What would a graph of pressure vs. volume of a gas look like according to Boyle's Law?

## Charles' Law: Answer the following questions about Charles' Law.

- Describe Charles' Law in your own words. **Increase temperature, increase volume**
- What is the formula?  **$V = kT$**
- What are the variables? What are the possible units used in the formula? **Temperature and volume (kelvin, mL, L)**
- What is constant? **Pressure**
- What is very important to remember about the temperature variable when using this formula? **It must be in kelvin**
- Give an example to help explain Charles's Law (you should have some type of visual of your example).



- What would a graph of temperature vs. volume of a gas look like according to Charles's Law?

## Gay-Lussac's Law: Answer the following questions about Gay-Lussac's Law.

- Describe Gay-Lussac's Law in your own words. **As pressure increases, temperature increases**
- What is the formula?  **$P = kT$**
- What are the variables? What are the possible units used in the formula? **Pressure and temperature (kelvin, atm, kPa, torr)**
- What is constant? **volume**
- What is very important to remember about the temperature variable when using this formula? **It must be in kelvin**
- Give an example to help explain Gay-Lussac's Law (you should have some type of visual of your example).

# Gas Laws Study Guide Answer Key

**Gary S. Thorpe**



## Gas Laws Study Guide Answer Key:

**Study Guide for Chemical Principles** Thomas Elliott Taylor,1979 Chemistry, Student Study Guide John A. Olmsted,Gregory M. Williams,2002-01-04 This third edition continues to innovate by providing students with an integrated and modern approach to the subject The text emphasizes the modern tools of chemistry while incorporating historical evidence and its unique molecular quantitative emphasis is further reinforced by an integrated media package developed by the authors Also of benefit is the just in time presentation of key content only providing details once they are needed While key topics and analytical techniques have been updated there is now an additional third chapter on chemical equilibrium The authors have also developed an expanded and more integrated problem solving emphasis that now incorporates a 4 step strategy throughout complete with text icons The whole is backed by a range of supplements including a new illustration program a tutorial CD interactive learningware an extensive Web CT component an instructor s resource CD and a solution CD *Study Guide, Introductory Chemistry, a Foundation ; Introductory Chemistry ; Basic Chemistry* Iris Stovall,Donald J. DeCoste,2000 Study Guide for the Telecourse Project Universe Joel M. Levine,Richard T. Searles,1988 **Study guide to accompany Drew H. Wolfe: General, organic and biological chemistry** Daniel C. Pantaleo,1986 *Chemistry* Gary S. Thorpe,2001 CliffsAP study guides help you gain an edge on Advanced Placement exams Review exercises realistic practice exams and effective test taking strategies are the key to calmer nerves and higher APa scores CliffsAP Chemistry is for students who are enrolled in AP Chemistry or who are preparing for the Advanced Placement Examination in Chemistry Inside you ll find hints for answering the essay and multiple choice sections a clear explanation of the exam format reviews of all 22 required labs a look at how exams are graded and more Realistic full length practice exam Answers to commonly asked questions about the AP Chemistry exam Study strategies to help you prepare Thorough review of the key topics that are sure to be on the test Sample laboratory write ups The AP Chemistry exam is coming up Your thorough understanding of months and months of college level chemistry coursework is about to be evaluated in a 3 hour examination CliffsAP Chemistry includes the following material to you do the very best job possible on the big test Gravimetrics Electronic structure of atoms Covalent bonding and ionic bonding Acids and bases Reduction and oxidation Organic chemistry and nuclear chemistry Writing and predicting chemical reactions This comprehensive guide offers a thorough review of key concepts and detailed answer explanations It s all you need to do your best and get the college credits you deserve a Advanced Placement Program and AP are registered trademarks of the College Board which was not involved in the production of and does not endorse this product *Study Guide to Accompany Human Physiology* Kevin T. Patton,Victoria Schauf,1989-12 **Study Guide for Project: Universe** Stephen P. Lattanzio,1981 **Prentice Hall Chemistry** ,2000 U.S. Naval Training Bulletin ,1948 Student Study Guide/solutions Manual to Accompany Foundations of Inorganic, Organic, and Biological Chemistry Katherine J. Denniston,1995 This student study guide has been written to accompany Foundations of Inorganic Organic and

Biological Chemistry by Caret Denniston and Topping It was designed to complement the text not to be used in place of the text Each chapter of the study guide contains the following sections reorganized and expanded set of learning goals concise chapter summary in chapter solved problems list of key terms review problems Preface *Chemistry* Zumdahl, Steven S. Zumdahl, 2003-07 Contains chapter discussions key term definitions and practice chapter tests **Introductory Chemistry** Steve Russo, Mike Silver, 2002 Steve Russo and Mike Silver turn chemistry into a memorable story that engages readers and provides the context they need to understand and remember core concepts The book builds interesting applications and well designed illustrations into the narrative to get and hold attention then builds confidence with integrated active learning activities Readers make the connections between concepts and the problem solving techniques they need to master as they read The new edition strengthens this conceptual approach and presents additional quantitative techniques in key areas Readers will find enhanced support for quantitative problem solving and more challenging questions at the end of each chapter in addition to the wealth of technology based support on The Chemistry Place tm Special Edition and on The Chemistry of Life CD ROM For college instructors and students Study Guide to Accompany Ekeland/Tollison Economics, Fourth Edition John Keith Watson, 1994 *Resources in Education* ,1976 *The Addison-Wesley Book of Apple Software 1984* Jeffrey Stanton, 1984 **Essentials of Introductory Chemistry** Russo Steve Silver Michael, Steve Russo, 2001-12 Introductory Chemistry Third Edition helps readers master the quantitative skills and conceptual understanding they need to gain a deep understanding of chemistry Unlike other books on the market that emphasize rote memory of problem solving algorithms Introductory Chemistry takes a conceptual approach with the idea that focusing on the concepts behind chemical equations helps readers become more proficient problem solvers What Is Chemistry The Numerical Side of Chemistry The Evolution of Atomic Theory The Modern Model of the Atom 1 Chemical Bonding and Nomenclature The Shape of Molecules Chemical Reactions Stoichiometry and the Mole The Transfer of Electrons from One Atom to Another in a Chemical Reaction Intermolecular Forces and the Phases of Matter What If There Were No Intermolecular Forces The Ideal Gas Solutions When Reactants Turn into Products Chemical Equilibrium Electrolytes Acids and Bases For all readers interested in introductory chemistry **Prentice Hall Physical Science Concepts in Action Program Planner National Chemistry Physics Earth Science** ,2003-11 Prentice Hall Physical Science Concepts in Action helps students make the important connection between the science they read and what they experience every day Relevant content lively explorations and a wealth of hands on activities take students understanding of science beyond the page and into the world around them Now includes even more technology tools and activities to support differentiated instruction Modules McDougal Littell Incorporated, 2005 **Catalogue of Title-entries of Books and Other Articles Entered in the Office of the Librarian of Congress, at Washington, Under the Copyright Law ... Wherein the Copyright Has Been Completed by the Deposit of Two Copies in the Office** Library of Congress. Copyright Office, 1943

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