

SHORT ANSWER QUESTIONS (From Previous Board Papers 2014-2015-2016-2017)

1. Why HF is weaker acid than HCl? (LHR, GH, 2015)

Ans. HF is weaker than HCl because ionization of HF is very small (10% only). Due to strong hydrogen bonding, HF has less ionization.

2. Give reason why fluorine is a gas, iodine is solid. (GRW, 2017)

Ans. Fluorine is a gas because it has small size, low polarizability and hence weak van der Waals forces, whereas iodine has large size, greater polarizability and strong London forces which make it solid.

3. In what ways fluorine differs from its group members? (BWP, SGB, GI, 2015)(MLN, GH, 2016)(SGD, GI, 2017)

Ans. **Peculiar Behaviour of Fluorine:** Fluorine shows peculiar behaviour in many respects which is due to:

- (i) Small size of F atom and of F^- ion. (ii) High first ionization energy and electronegativity.
- (iii) Low dissociation energy of F_2 molecule as compared to Cl_2 and Br_2 .
- (iv) Restriction of valence shell to an octet. (v) Direct combination with inert gases.

4. Describe two uses of helium. (LHR, GI, 2016)

Ans. Helium is first member of group VIII A, and found as a mono atomic gas in nature. Helium has following important uses.

(i) **In weather balloons:** Helium is used in weather balloons, in welding and in traffic signal light.

(ii) **In sea diving:** A mixture of 80% helium and 20% oxygen is used for breathing by the sea divers.

(iii) **In nuclear reactors:** Helium is used as a cooling medium for nuclear reactors.

5. Halogens are strong oxidizing agents justify. (SWL, MLN, GH, 2014) (MLN, GI, 2016)

Ans. All halogens have positive values of reduction potential, they have tendency to accept electron and form negative ion. They can remove electron from other elements and oxidize them to positive ions, e.g. $2Na + F_2 \longrightarrow 2NaF$

Hence it is justified that halogens are oxidizing agents. The order of oxidizing power of halogens is as follows: $F_2 > Cl_2 > Br_2 > I_2$

Gases Section Short Answers

David Baud



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Table of Contents Gases Section Short Answers

1. Understanding the eBook Gases Section Short Answers
 - The Rise of Digital Reading Gases Section Short Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Gases Section Short Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gases Section Short Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gases Section Short Answers
 - Personalized Recommendations
 - Gases Section Short Answers User Reviews and Ratings
 - Gases Section Short Answers and Bestseller Lists

5. Accessing Gases Section Short Answers Free and Paid eBooks
 - Gases Section Short Answers Public Domain eBooks
 - Gases Section Short Answers eBook Subscription Services
 - Gases Section Short Answers Budget-Friendly Options
6. Navigating Gases Section Short Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Gases Section Short Answers Compatibility with Devices
 - Gases Section Short Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Gases Section Short Answers
 - Highlighting and Note-Taking Gases Section Short Answers
 - Interactive Elements Gases Section Short Answers
8. Staying Engaged with Gases Section Short Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gases Section Short Answers
9. Balancing eBooks and Physical Books Gases Section Short Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gases Section Short Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Gases Section Short Answers
 - Setting Reading Goals Gases Section Short Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Gases Section Short Answers
 - Fact-Checking eBook Content of Gases Section Short Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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