

Gas Stoichiometry problems

1. Calculate the mass of hydrogen peroxide needed to obtain 0.460 L of oxygen gas at STP.



$$0.460 \text{ L O}_2 \left| \frac{1 \text{ mole O}_2}{22.4 \text{ L O}_2} \right| \left| \frac{2 \text{ moles H}_2\text{O}_2}{1 \text{ mole O}_2} \right| \left| \frac{34.0 \text{ g H}_2\text{O}_2}{1 \text{ mole H}_2\text{O}_2} \right| = \boxed{1.40 \text{ g}}$$

Given:

$V = 0.460 \text{ L}$   
 $T = 273 \text{ K}$   
 $P = 1 \text{ atm}$   
 $n = ?$

2. Sodium azide ( $\text{NaN}_3$ ) is used in some air bags. The impact of a collision causes the sodium azide to decompose and form solid Sodium and nitrogen gas. Calculate the volume of nitrogen gas generated at  $80.0^\circ\text{C}$  and  $823 \text{ mm Hg}$  by the decomposition of  $60.0 \text{ g}$  of  $\text{NaN}_3$ .

Balanced equation:



Given:

$T = 353 \text{ K}$   
 $P = 823 \text{ mm Hg}$   
 $V_{\text{N}_2} = ?$

$$60.0 \text{ g NaN}_3 \left| \frac{1 \text{ mole NaN}_3}{65.0 \text{ g NaN}_3} \right| \left| \frac{3 \text{ moles N}_2}{2 \text{ moles NaN}_3} \right| = 1.38 \text{ moles N}_2$$

$$V_{\text{N}_2} = \frac{nRT}{P} = \frac{(1.38 \text{ moles}) (0.0821 \frac{\text{L atm}}{\text{mol K}}) (353 \text{ K})}{1.08 \text{ atm}}$$

$$V = 37.0 \text{ L}$$

$$823 \text{ mm Hg} \left| \frac{1 \text{ atm}}{760 \text{ mm Hg}} \right| = 1.08 \text{ atm}$$

3. When heated, solid potassium chlorate ( $\text{KClO}_3$ ) decomposes to form solid potassium chloride and oxygen gas. If  $20.8 \text{ g}$  of potassium chlorate decomposes, how many liters of oxygen gas will form at STP?

Balanced equation:



Given:

$m = 20.8 \text{ g}$   
 $T = 273 \text{ K}$   
 $P = 1 \text{ atm}$   
 $V = ?$   
 $\boxed{\text{STP}}$

$$20.8 \text{ g KClO}_3 \left| \frac{1 \text{ mole KClO}_3}{122.55 \text{ g KClO}_3} \right| \left| \frac{3 \text{ moles O}_2}{2 \text{ moles KClO}_3} \right| = 0.801 \text{ moles O}_2$$

$$V_{\text{O}_2} = \frac{nRT}{P} = \frac{(0.801 \text{ moles}) (0.0821 \frac{\text{L atm}}{\text{mol K}}) (273 \text{ K})}{1 \text{ atm}}$$

$$\boxed{V_{\text{O}_2} = 5.72 \text{ Liters}}$$

OR - Since at STP - do in one step.

$$20.8 \text{ g KClO}_3 \left| \frac{1 \text{ mole KClO}_3}{122.55 \text{ g KClO}_3} \right| \left| \frac{3 \text{ moles O}_2}{2 \text{ moles KClO}_3} \right| \left| \frac{22.4 \text{ L O}_2}{1 \text{ mole O}_2} \right| = \boxed{5.70 \text{ L}}$$

# Gas Stoichiometry Worksheet Answers With Work

**DJ Losen**



## **Gas Stoichiometry Worksheet Answers With Work:**

## Decoding **Gas Stoichiometry Worksheet Answers With Work**: Revealing the Captivating Potential of Verbal Expression

In an era characterized by interconnectedness and an insatiable thirst for knowledge, the captivating potential of verbal expression has emerged as a formidable force. Its capability to evoke sentiments, stimulate introspection, and incite profound transformations is genuinely awe-inspiring. Within the pages of "**Gas Stoichiometry Worksheet Answers With Work**," a mesmerizing literary creation penned with a celebrated wordsmith, readers set about an enlightening odyssey, unraveling the intricate significance of language and its enduring affect our lives. In this appraisal, we shall explore the book is central themes, evaluate its distinctive writing style, and gauge its pervasive influence on the hearts and minds of its readership.

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