

### Gas Stoichiometry Practice KEY

Solve for the unknown masses and volumes using gas laws and stoichiometry

1. If you have 2 L of A and plenty of B then how many moles of C can be produced at STP?



$$\frac{2 \text{ L A}}{1} \times \frac{1 \text{ mol}}{22.4 \text{ L}} \times \frac{2 \text{ mol C}}{2 \text{ mol A}} = 0.089 \text{ mol C}$$

2. 10L of hydrogen gas with plenty of ammonia at STP will produce what volume of ammonium?



$$\frac{10 \text{ L H}_2}{1} \times \frac{1 \text{ mol}}{22.4 \text{ L}} \times \frac{2 \text{ mol NH}_4}{1 \text{ mol H}_2} \times 22.4 \text{ L} = 20 \text{ L NH}_4$$

3. 5L of hydrogen gas and 4L of oxygen gas at STP will produce what mass of water?



$$5 \text{ L} / 22.4 \text{ L} = .223 \text{ mol H}_2$$

$$4 \text{ L} / 22.4 \text{ L} = .179 \text{ mol O}_2$$

$$\frac{.223 \text{ mol H}_2}{1} \times \frac{1 \text{ mol O}_2}{2 \text{ mol H}_2} = .111 \text{ mol O}_2 \quad .111 \text{ mol O}_2 \text{ needed} < .179 \text{ mol O}_2 \text{ have. H}_2 \text{ is LR}$$

$$\frac{.223 \text{ mol H}_2}{1} \times \frac{2 \text{ mol H}_2O}{2 \text{ mol H}_2} \times \frac{18 \text{ g H}_2O}{1 \text{ mol H}_2O} = 4.01 \text{ g H}_2O$$

4. 100g of carbon and 3L of oxygen gas under the conditions of 5 atm and 50°C should produce how much carbon dioxide gas (volume)?



$$100 \text{ g} / 12 \text{ g} = 8.33 \text{ mol C}$$

$$5(3) = n(.0821)323 \Rightarrow .566 \text{ mol O}_2$$

$$\frac{8.33 \text{ mol C}}{1} \times \frac{1 \text{ mol O}_2}{1 \text{ mol C}} = 8.33 \text{ mol O}_2$$

$$8.33 \text{ mol O}_2 \text{ needed} > .566 \text{ mol O}_2 \text{ have. O}_2 \text{ is LR}$$

$$\frac{0.566 \text{ mol O}_2}{1} \times \frac{1 \text{ mol CO}_2}{1 \text{ mol O}_2} = 0.566 \text{ mol CO}_2$$

$$5(v) = .566(.0821)323 \Rightarrow 3 \text{ L CO}_2$$

5. 20L of methane and excess oxygen gas will produce what volume of carbon dioxide and mass of water if the reactants are under 20 atm and 0°C?

$$20(20) = n(.0821)273 \Rightarrow 17.85 \text{ mol CH}_4$$

# Gas Stoichiometry Practice Answer

**JR Anderson**



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## **Table of Contents Gas Stoichiometry Practice Answer**

1. Understanding the eBook Gas Stoichiometry Practice Answer
  - The Rise of Digital Reading Gas Stoichiometry Practice Answer
  - Advantages of eBooks Over Traditional Books
2. Identifying Gas Stoichiometry Practice Answer
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Gas Stoichiometry Practice Answer
  - User-Friendly Interface
4. Exploring eBook Recommendations from Gas Stoichiometry Practice Answer
  - Personalized Recommendations
  - Gas Stoichiometry Practice Answer User Reviews and Ratings
  - Gas Stoichiometry Practice Answer and Bestseller Lists

5. Accessing Gas Stoichiometry Practice Answer Free and Paid eBooks
  - Gas Stoichiometry Practice Answer Public Domain eBooks
  - Gas Stoichiometry Practice Answer eBook Subscription Services
  - Gas Stoichiometry Practice Answer Budget-Friendly Options
6. Navigating Gas Stoichiometry Practice Answer eBook Formats
  - ePub, PDF, MOBI, and More
  - Gas Stoichiometry Practice Answer Compatibility with Devices
  - Gas Stoichiometry Practice Answer Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Gas Stoichiometry Practice Answer
  - Highlighting and Note-Taking Gas Stoichiometry Practice Answer
  - Interactive Elements Gas Stoichiometry Practice Answer
8. Staying Engaged with Gas Stoichiometry Practice Answer
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Gas Stoichiometry Practice Answer
9. Balancing eBooks and Physical Books Gas Stoichiometry Practice Answer
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Gas Stoichiometry Practice Answer
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Gas Stoichiometry Practice Answer
  - Setting Reading Goals Gas Stoichiometry Practice Answer
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Gas Stoichiometry Practice Answer
  - Fact-Checking eBook Content of Gas Stoichiometry Practice Answer
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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